A reaction in which the <u>Products</u> can react to form the <u>reactants</u> is called a <u>reversible</u> reaction.

Chemical equilibrium occurs when the reactants in a reversible reaction form products at the same rate that products form

At equilibrium:

- the <u>concentration</u> of the reactants and products does not change
- the concentration of reactants can be __equal __to, __greater_, or ___less___ the concentration of the products.

$$K^{*a} = \frac{[],[],[],}{[],[],[],}$$

If Keg is	reaction is favored
= 1	neither
< 1	reverse
> 1	forward

Ex. #1: Calculate K_{eq} for the following reaction using the given equilibrium concentrations. Then determine whether the forward or reverse reaction is favored.

$$N_2O_4 \implies 2 NO_2[N_2O_4] = 0.23M [NO_2] = 0.037M$$

Note Taking Guide Episode 1202 Answers

Brendan G. Carr

Note Taking Guide Episode 1202 Answers:

The Wireless World ,1933 New York Times Saturday Book Review Supplement ,1993-04

This book delves into Note Taking Guide Episode 1202 Answers. Note Taking Guide Episode 1202 Answers is a crucial topic that needs to be grasped by everyone, ranging from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Note Taking Guide Episode 1202 Answers, encompassing both the fundamentals and more intricate discussions.

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 - o Chapter 2: Essential Elements of Note Taking Guide Episode 1202 Answers
 - Chapter 3: Note Taking Guide Episode 1202 Answers in Everyday Life
 - Chapter 4: Note Taking Guide Episode 1202 Answers in Specific Contexts
 - \circ Chapter 5: Conclusion
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