

Questions with Solution P. K. Nag

- Q4.1** An engine is tested by means of a water brake at 1000 rpm. The measured torque of the engine is 10000 mN and the water consumption of the brake is 0.5 m³/s, its inlet temperature being 20°C. Calculate the water temperature at exit, assuming that the whole of the engine power is ultimately transformed into heat which is absorbed by the cooling water.

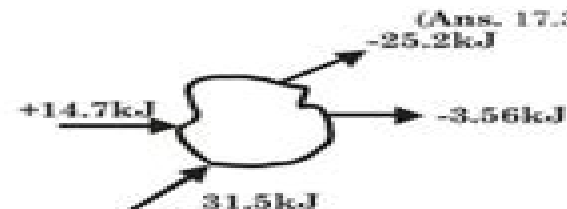
(Ans. 20.5°C)

Solution: Power = $T\omega$
 $= 10000 \times \left(\frac{2\pi \times 1000}{60} \right)$
 $= 1.0472 \times 10^6 \text{ W}$
 $= 1.0472 \text{ MW}$
 Let final temperature = $t^\circ\text{C}$
 $\therefore$ Heat absorb by cooling water / unit = $\dot{m} \times \Delta t$
 $= \dot{V} \rho \times \Delta t$
 $= 0.5 \times 1000 \times 4.2 \times (t - 20)$
 $\therefore 0.5 \times 1000 \times 4.2 \times (t - 20) = 1.0472 \times 10^6$
 $\therefore t - 20 = 0.4986 \approx 0.5$
 $\therefore t = 20.5^\circ\text{C}$

- Q4.2** In a cyclic process, heat transfers are + 14.7 kJ, - 25.2 kJ, - 3.56 kJ and + 31.5 kJ. What is the net work for this cyclic process?

(Ans. 17.34 kJ)

Solution : $\sum Q = (14.7 + 31.5 - 25.2 - 3.56) \text{ kJ}$
 $= 17.44 \text{ kJ}$
 From first law of thermodynamics
 (for a cyclic process)
 $\sum Q = \sum W$
 $\therefore \sum W = 17.44 \text{ kJ}$



- Q4.3** A slow chemical reaction takes place in a fluid at the constant pressure of 0.1 MPa. The fluid is surrounded by a perfect heat insulator during the reaction which begins at state 1 and ends at state 2. The insulation is then removed and 105 kJ of heat flow to the surroundings as the fluid goes to state 3. The following data are observed for the fluid at states 1, 2 and 3.

State	$v \text{ (m}^3\text{)}$	$t \text{ (}^\circ\text{C)}$
1	0.003	20
2	0.3	370
3	0.06	20

For the fluid system, calculate E_2 and E_3 , if $E_1 = 0$

(Ans. $E_2 = -29.7 \text{ kJ}$, $E_3 = -110.7 \text{ kJ}$)

Pk Nag Solution Manual

Lingsheng Yao



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