#### Personalised Learning Checklists AQA Chemistry Paper 1



Topi	Student Checklist	R	A	6
4.1.1 A simple model of the atom, symbols, relative atomic mass, electronic charge and isotopes	State that everything is made of atoms and recall what they are			_
	Describe what elements and compounds are			_
	State that elements and compounds are represented by symbols, and use chemical symbols and			т
	formulae to represent elements and compounds			1
	Write word equations and balanced symbol equations for chemical reactions, including using			П
	appropriate state symbols			
	HT ONLY: Write balanced half equations and ionic equations			
	Describe what a mixture is		3.5	
	Name and describe the physical processes used to separate mixtures and suggest suitable separation			П
	techniques			L
	Describe how the atomic model has changed over time due to new experimental evidence, inc discovery			П
	of the atom and scattering experiments (inc the work of James Chadwick)			L
	Describe the difference between the plum pudding model of the atom and the nuclear model of the			П
	atom			
	State the relative charge of protons, neutrons and electrons and describe the overall charge of an atom			
	State the relative masses of protons, neutrons and electrons and describe the distribution of mass in an			ш
	atom	_		┺
	Calculate the number of protons, neutrons and electrons in an atom when given its atomic number and	1 1		1
	mass number  Describe isotopes as atoms of the same element with different numbers of neutrons	-		₽
	Describe isotopes as atoms of the same element with different numbers of neutrons  Define the term relative atomic mass and why it takes into account the abundance of isotopes of the	-	-	₽
	Define the term relative atomic mass and why it takes into account the abundance of isotopes of the			1
	Calculate the relative atomic mass of an element given the percentage abundance of its isotopes	-	_	₽
	Describe how electrons till energy levels in atoms, and represent the electron structure of elements	-	-	۰
	using discrems and numbers			ш
	Recall how the elements in the periodic table are arranged	-		t
4.1.2 The periodic table	Describe how elements with similar properties are placed in the periodic table	_		t
	Explain why elements in the same group have similar properties and how to use the periodic table to	_	_	t
	oredict the reactivity of elements			ш
	Describe the early attempts to classify elements	_		t
	Explain the creation and attributes of Mendeleev's periodic table			t
	identify metals and non-metals on the periodic table, compare and contrast their properties	-		t
	Explain how the atomic structure of metals and non-metals relates to their position in the periodic table			m
	Describe nobel gases (group 0) and explain their lack of reactivity			т
	Describe the properties of noble gases, including boiling points, predict trends down the group and			Т
	describe how their properties depend on the outer shell of electrons			ш
	Describe the reactivity and properties of group 1 alkali metals with reference to their electron arrangement and predict their reactions			Γ
	Describe the properties of group 7 halogens and how their properties relate to their electron			П
	arrangement, including trends in molecular mass, melting and boiling points and reactivity			
	Describe the reactions of group 7 halogens with metals and non-metals			Г
	Chem ONLY: Describe the properties of transition metals and compare them with group I elements, including melting points and densities, strength and hardness, and reactivity (for CR, Mn Fe, Co, Ni & Cu)			

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