

Annotated periodic table of the elements

1 H1.008Hydrogen																		13 B10.81Boron	14 C12.01Carbon	15 N14.01Nitrogen	16 O16.00Oxygen	17 F18.99Fluorine	18 Ne20.18Neon																									
3 Li6.94Lithium		4 Be9.012Beryllium																	19 Al26.98Aluminium	20 Si28.09Silicon	21 P30.97Phosphorus	22 S32.06Sulfur	23 Cl35.45Chlorine	24 Ar39.95Argon																								
11 Na22.99Sodium		12 Mg24.31Magnesium																	25 Sc44.96Scandium	26 Ti47.87Titanium	27 V50.94Vanadium	28 Cr52.00Chromium	29 Mn54.94Manganese	30 Fe55.85Iron	31 Co58.93Cobalt	32 Ni58.69Nickel	33 Cu63.55Copper	34 Zn65.38Zinc	35 Ga69.72Gallium	36 Ge72.64Germanium	37 As74.92Arsenic	38 Se78.96Selenium	39 Br79.90Bromine	40 Kr83.80Krypton														
19 K39.10Potassium		20 Ca40.08Calcium																	41 Sc44.96Scandium	42 Ti47.87Titanium	43 V50.94Vanadium	44 Cr52.00Chromium	45 Mn54.94Manganese	46 Fe55.85Iron	47 Co58.93Cobalt	48 Ni58.69Nickel	49 Cu63.55Copper	50 Zn65.38Zinc	51 Ga69.72Gallium	52 Ge72.64Germanium	53 As74.92Arsenic	54 Se78.96Selenium	55 Br79.90Bromine	56 Kr83.80Krypton														
37 Rb85.47Rubidium		38 Sr87.62Strontium																	57 La138.91Lanthanum	58 Ce140.12Cerium	59 Pr140.91Praseodymium	60 Nd144.24Neodymium	61 Pm144.91Promethium	62 Sm150.36Samarium	63 Eu151.96Europium	64 Gd157.25Gadolinium	65 Tb158.93Terbium	66 Dy162.50Dysprosium	67 Ho164.93Holmium	68 Er167.26Erbium	69 Tm168.93Thulium	70 Yb173.05Ytterbium	71 Lu174.97Lutetium	72 Hf178.49Hafnium	73 Ta180.95Tantalum	74 W183.84Tungsten	75 Re186.21Rhenium	76 Os190.23Osmium	77 Ir192.22Iridium	78 Pt195.08Platinum	79 Au196.97Gold	80 Hg200.59Mercury	81 Tl204.38Thallium	82 Pb207.2Lead	83 Bi208.98Bismuth	84 Po209Polonium	85 At210Astatine	86 Rn222Radon
55 Cs132.91Cesium		56 Ba137.33Barium																	89 La138.91Lanthanum	90 Ce140.12Cerium	91 Pr140.91Praseodymium	92 Nd144.24Neodymium	93 Pm144.91Promethium	94 Sm150.36Samarium	95 Eu151.96Europium	96 Gd157.25Gadolinium	97 Tb158.93Terbium	98 Dy162.50Dysprosium	99 Ho164.93Holmium	100 Er167.26Erbium	101 Tm168.93Thulium	102 Yb173.05Ytterbium	103 Lu174.97Lutetium	104 Hf178.49Hafnium	105 Ta180.95Tantalum	106 W183.84Tungsten	107 Re186.21Rhenium	108 Os190.23Osmium	109 Ir192.22Iridium	110 Pt195.08Platinum	111 Au196.97Gold	112 Hg200.59Mercury	113 Tl204.38Thallium	114 Pb207.2Lead	115 Bi208.98Bismuth	116 Po209Polonium	117 At210Astatine	118 Rn222Radon
87 Fr223Francium		88 Ra226Radium																	103 La138.91Lanthanum	104 Rf181Rutherfordium	105 Db185Dubnium	106 Sg271Seaborgium	107 Bh270Bohrium	108 Hs277Hassium	109 Mt276Meitnerium	110 Ds285Darmstadtium	111 Rg284Roentgenium	112 Cn285Copernicium	113 Nh286Nihonium	114 Fl289Flerovium	115 Mc288Moscovium	116 Lv293Livermorium	117 Ts294Tennessine	118 Og294Oganesson														

Group names*

- Hydrogen & the alkali metals
- Alkaline earth metals
- Coinage metals
- Volatile metals
- Halogens
- Chalcogens
- Pnictogens
- Chalcogens
- Halogens
- Helium group

*Groups 3 to 10 are named after their first members i.e. Group 3 is the Scandium group

Main group
oxidation states
electronegativity
atomic radius

Atomic number
and symbol
Atomic weight
and name

CLASS COUNTERPARTS

Metals	Nonmetals
Pre-transition	
Active	Active
Transition	Intermediate
Post-transition	Metals
Metals (Ru-Pt, Os-Au)	Metals

Active metals

Pre-transition metals
Light, reactive, mostly of low mechanical strength, melting and boiling points. The group 1 members are easily cut with a knife. Only Be, Mg and Al have any structural uses. Most form basic oxides (Be and Al are amphoteric).

Rare earths
Closely resemble Ca, Sr and Ba, but many are heavier and they form mostly pale coloured compounds. Similar to one another, esp. La to Lu (aka the lanthanides), and hard to separate for that reason. Abundant but widely dispersed; commercially viable concentrations are rare. Oxides are strongly to moderately basic. Sc, Yb and Lu exhibit weak amphoteric.

Actinides
Soft, dense and reactive metals. Those from Ra onwards are synthetic. Early actinides (Th-Am) show some similarities to transition metals and have basic, amphoteric, or acidic oxides; late actinides are more like lanthanides. Many actinides form coloured compounds.

Transition metals

Moderate to high density metals, with high melting and boiling points, many have high hardness and mechanical strength. Those from Rf onwards are synthetic. Chemically show variable valency and strong tendency to form coordination and colored compounds. Oxides are basic, amphoteric or acidic, depending on oxidation state.

Noble metals
Resistant to corrosion and oxidation, unlike base metals such as Fe, Ni, Cu and Zn. Commonly found together in native form. Good catalytic properties. Very weakly electropositive; insoluble oxides.

Frontier metals
Soft (or brittle), low strength, and with melting points lower than transition metals. Crystalline structures tend to show directional bonding, with generally greater complexity or lower nearest neighbours than other metals. All used or tried as soft metal lubricants. Most are chemically weak; a minority are moderately electropositive. Characterised, to varying degrees, by covalent bonding tendencies, amphoteric, and anionic species such as zincates, stannates, and bismuthates.

Metals

Metals
Metallic looking, brittle, fair electrical conductors. All known in semiconducting forms. Moderate ionisation energies (750–1,000 kJ/mol), and electronegativity values (1.8–2.3). Mostly behave chemically as non-metals. Can form alloys with metals. Amphoteric or weakly acidic oxides.

Intermediate nonmetals
H and N are colorless gases; C, P, and Se are metalloids looking solids while S is yellow. The latter four are brittle but all known in malleable, plastic or ductile forms. Moderate to high ionisation energies, low to high electron affinities, moderate to high electronegativity values, relatively poor to moderately strong oxidising agents, tend to form covalent or metallic to near-metallic compounds with metals. Their oxides are acidic, or neutral (H₂O, CO, NO, N₂O).

*as liquid P, the most stable form in ambient conditions.

Group 1

This group is Sc-Y, and the lanthanides (La-Lu) and the actinides (Ac-Lu). Trends going from Sc-Y and the lanthanides (La-Lu) trends going from Sc-Y and the actinides (Ac-Lu) trends going from Sc-Y and the actinides (Ac-Lu).

Nonmetals

Chalcogens
O, F, and Cl are, respectively, colorless, pale yellow, and pale yellow-green gases. Br is a red-brown liquid. I is a metallic looking solid. Chemically tend to have to high ionisation energies, electron affinities, and electronegativity values, and be relatively strong oxidising agents. Manifestations include oxygen's major association with the ubiquitous processes of corrosion and combustion, and the intrinsically corrosive nature of the non-metallic halogens. All five tend to form predominantly ionic compounds with metals. Their oxides are acidic.

Noble gases
Colorless, odorless, non-flammable gases with very low chemical reactivity. The first noble gas compound, of approximate composition XePF₆, was not prepared until 1962. More than half-a-thousand noble gas compounds were synthesised in the ensuing half-century. Compounds of He, Ne, and Ar, in ambient conditions, are not known.

Radioactive

Oganesson is likely to be either a metalloid or a post-transition metal.

Periodic Table Guide Answer Sheet

Lawrie Ryan

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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