

Annotated periodic table of the elements

Group names*

- 1 Hydrogen & the alkali metals
- 2 Alkaline earth metals
- 3 Coinage metals
- 12 Volatile metals
- 13 Borogens
- 14 Chalcogens
- 15 Pnictogens
- 16 Chalcogens
- 17 Halogens
- 18 Helium group

*Groups 3 to 10 are named after their first members i.e. Group 3 is the Scandium group

1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18
1 H 1.008 Hydrogen												B 10.81 Boron	C 12.01 Carbon	N 14.01 Nitrogen	O 16.00 Oxygen	F 18.99 Fluorine	Ne 20.18 Neon
3 Li 6.94 Lithium	4 Be 9.012 Beryllium											Al 26.98 Aluminum	Si 28.09 Silicon	P 30.97 Phosphorus	S 32.06 Sulfur	Cl 35.45 Chlorine	Ar 39.95 Argon
5 Na 22.99 Sodium	6 Mg 24.31 Magnesium											K 39.10 Potassium	Ca 40.08 Calcium				Kr 83.80 Krypton
7 K 39.10 Potassium	8 Ca 40.08 Calcium											Rb 85.47 Rubidium	Sr 87.62 Strontium				Xe 131.29 Xenon
9 Rb 85.47 Rubidium	10 Sr 87.62 Strontium											Cs 132.91 Cesium	Ba 137.33 Barium				Rn 222.02 Radon
11 Cs 132.91 Cesium	12 Ba 137.33 Barium											Fr 223.02 Francium	Ra 226.03 Radium				
13 Pr 140.91 Praseodymium	14 Nd 144.24 Neodymium	15 Pm 144.91 Promethium	16 Sm 150.36 Samarium	17 Eu 151.96 Europium	18 Gd 157.25 Gadolinium	19 Tb 158.93 Terbium	20 Dy 162.50 Dysprosium	21 Ho 164.93 Holmium	22 Er 167.26 Erbium	23 Tm 168.93 Thulium	24 Yb 173.05 Ytterbium						
25 Mn 54.94 Manganese	26 Fe 55.85 Iron	27 Co 58.93 Cobalt	28 Ni 58.69 Nickel	29 Cu 63.55 Copper	30 Zn 65.38 Zinc	31 Ga 69.72 Gallium	32 Ge 72.64 Germanium	33 As 74.92 Arsenic	34 Se 78.96 Selenium	35 Br 79.90 Bromine	36 Kr 83.80 Krypton						
37 Y 88.91 Yttrium	38 Zr 91.22 Zirconium	39 Nb 92.91 Niobium	40 Mo 95.94 Molybdenum	41 Tc 98.91 Technetium	42 Ru 101.07 Ruthenium	43 Rh 102.91 Rhodium	44 Pd 106.42 Palladium	45 Ag 107.87 Silver	46 Cd 112.41 Cadmium	47 In 114.82 Indium	48 Sn 118.71 Tin	49 Sb 121.76 Antimony	50 Te 127.60 Tellurium	51 I 126.91 Iodine	52 Xe 131.29 Xenon		
53 La 138.91 Lanthanum	54 Ce 140.12 Cerium	55 Pr 140.91 Praseodymium	56 Nd 144.24 Neodymium	57 Pm 144.91 Promethium	58 Sm 150.36 Samarium	59 Eu 151.96 Europium	60 Gd 157.25 Gadolinium	61 Tb 158.93 Terbium	62 Dy 162.50 Dysprosium	63 Ho 164.93 Holmium	64 Er 167.26 Erbium	65 Tm 168.93 Thulium	66 Yb 173.05 Ytterbium				
67 Lu 174.97 Lutetium	68 Hf 178.49 Hafnium	69 Ta 180.95 Tantalum	70 W 183.84 Tungsten	71 Re 186.21 Rhenium	72 Os 190.23 Osmium	73 Ir 192.22 Iridium	74 Pt 195.08 Platinum	75 Au 196.97 Gold	76 Hg 200.59 Mercury	77 Tl 204.38 Thallium	78 Pb 207.2 Lead	79 Bi 208.98 Bismuth	80 Po 209 Polonium	81 At 210 Astatine	82 Rn 222 Radon		
83 Fr 223.02 Francium	84 Ra 226.03 Radium																
85 At 210 Astatine	86 Rn 222 Radon																

CLASSIFICATION

Metals	Nonmetals
Pre-transition	
Active	Post-transition
Alkaline earth	
Alkaloids	
Transition	Intermediate
Post-transition	
Metals	
Metals (Ru-Pt, Os-Au)	Metals gas

87 La 138.91 Lanthanum	88 Ce 140.12 Cerium	89 Pr 140.91 Praseodymium	90 Nd 144.24 Neodymium	91 Pm 144.91 Promethium	92 Sm 150.36 Samarium	93 Eu 151.96 Europium	94 Gd 157.25 Gadolinium	95 Tb 158.93 Terbium	96 Dy 162.50 Dysprosium	97 Ho 164.93 Holmium	98 Er 167.26 Erbium	99 Tm 168.93 Thulium	100 Yb 173.05 Ytterbium
101 Ac 227.03 Actinium	102 Th 232.04 Thorium	103 Pa 231.04 Protactinium	104 U 238.03 Uranium	105 Np 237.05 Neptunium	106 Pu 244.06 Plutonium	107 Am 243.06 Americium	108 Cm 247.07 Curium	109 Bk 247.07 Berkelium	110 Cf 251.08 Californium	111 Es 252.08 Einsteinium	112 Fm 257.10 Fermium	113 Md 258.10 Mendelevium	114 No 259.10 Nobelium

Radioactive

Approximate
likely to be either
a metalloid or a
post-transition
metal

Metals

Transition metals

Moderate to high density metals, with high melting and boiling points, many have high hardness and mechanical strength. Those from Rf onwards are synthetic. Chemically show variable valency and strong tendency to form coordination and colored compounds. Oxides are basic, amphoteric or acidic, depending on oxidation state.

Noble metals

Resistant to corrosion and oxidation, unlike base metals such as Fe, Ni, Cu and Zn. Commonly found together in native form. Good catalyst properties. Very weakly electropositive, inert to acids.

Post-transition

Soft (or brittle), low strength, and with melting points lower than transition metals. Crystalline structures tend to show directional bonding, with generally greater complexity or lower nearest neighbours than other metals. All used or tried as soft metal lubricants. Most are chemically weak; a minority are moderately electropositive. Characterized, to varying degrees, by covalent bonding tendencies, amphoteric, and amionic species such as zincates, stannates, and bismuthates.

Metalloids

Metallic looking brittle, fair electrical conductors. All known in semiconducting forms. Moderate ionization energies (750–1,000 kJ/mol), and electronegativity values (1.8–2.3). Mostly behave chemically as non-metals. Can form alloys with metals. Amphoteric or weakly acidic oxides.

Intermediate nonmetals

H and N are colorless gases, C, P, and Se are metalloids looking solids while Si is yellow. The latter four are brittle but all known in malleable, plastic or ductile forms. Moderate to high ionization energies, low to high electron affinities, moderate to high electronegativity values, relatively poor to moderately strong oxidizing agents, tend to form covalent or metallic to near-metallic compounds with metals. Their oxides are acidic, or neutral (P₂O, CO, NO, N₂O).

*as black P, the most stable form in ambient conditions

Group 3

This starts as Sc, Ti, and then bifurcates into a La branch i.e. La-Ac, and a Lu branch i.e. Lu-Lf. Trends going down Sc-Ti, and then La branch, resemble those of groups 1–2.

• Lu branch, more resemble those of groups 1–2

Corrosive nonmetals

O, F, and Cl are, respectively, colorless, pale yellow, and pale yellow-green gases. Br is a red-brown liquid, I is a metallic looking solid. Chemically tend to have to high ionization energies, electron affinities, and electronegativity values, and be relatively strong oxidizing agents. Manifestations include oxygen's major association with the ubiquitous processes of corrosion and combustion, and the intrinsically corrosive nature of the non-metallic halogens. All five tend to form predominantly ionic compounds with metals. Their oxides are acidic.

Noble gases

Colorless, colorless, non-flammable gases with very low chemical reactivity. The first noble gas compound, of approximate composition XePF₆, was not prepared until 1962. More than half-a-thousand noble gas compounds were synthesized in the ensuing half-century. Compounds of He, Ne, and Ar, in ambient conditions, are not known.

Periodic Table Guide Answers

Team Prabhat

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located to the right of the 'Team Prabhat' text.

Periodic Table Guide Answers:

The Practice of Chemistry Study Guide & Solutions Manual Pamela Mills, Amina El-Ashmawy, 2003-04-14 Designed to help students understand the material better and avoid common mistakes Also includes solutions and explanations to odd numbered exercises

New Pattern NTA JEE Main Quick Guide in Chemistry with Numeric Answer Questions 3rd Edition Disha Experts, 2019-10-01 As NTA introduces Numeric Answer Questions in JEE Main Disha launches the Questions the 3rd latest updated edition of New Pattern NTA JEE Main Quick Guide in Chemistry with Numeric Answer Questions This study material is developed for quick revision and practice of the complete syllabus of the JEE Main Exam in a short span of 40 days The book can prove to be the ideal material for class 12 students as they can utilise this book to revise their preparation immediately after the board exams The book contains 27 chapters of class 11 12 and each Chapter contains JEE Main 6 Years at a Glance i e JEE Main 2019 2014 with TOPIC WISE Analysis Detailed Concept Maps covers entire JEE Syllabus for speedy revision IMPORTANT CRITICAL Points of the Chapter for last minute revision TIPS to PROBLEM SOLVING to help students to solve Problems in shortest possible time Exercise 1 CONCEPT BUILDER A Collection of Important Topic wise MCQs to Build Your Concepts Exercise 2 CONCEPT APPLICATOR A Collection of Quality MCQs that helps sharpen your concept application ability Exercise 3 Numeric Answer Questions A Collection of Quality Numeric Answer Questions as per the new pattern of JEE Answer Keys Detailed Solutions of all the Exercises and Past years problems are provided at the end of the chapter

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