

Study Guide – Unit 1 Periodic Table ANSWERS

Structure of an atom

- List the three subatomic particles which make up an atom. Include their location, charge, and atomic mass.

a. Electron	Negative charge	Electron Cloud
b. Proton	Positive charge	Nucleus
c. Neutron	Neutral charge	Nucleus
- In a neutral atom, which two subatomic particles are equal in number to each other?
Protons and Electrons
- What does the atomic number of an element indicate?
The number of protons in the atom of that element
- What is an isotope?
Atoms with the same number of protons, but different masses due to a different number of neutrons
- Will two different atoms of carbon always have the same number of protons?
Yes

Element key

13
Al
26.981
Aluminum

- What are the number of protons, neutrons, and electrons in an atom of aluminum?

a. Protons = 13
b. Neutrons = 14
c. Electrons = 13

Bohr Model

- How do you know how many energy levels (orbits) a Bohr model will have?
By the period the element is located in
- How many electrons can each energy level hold?

$1^{\text{st}} = 2$	$2^{\text{nd}} = 8$	$3^{\text{rd}} = 18$
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The Periodic Table

- How is the periodic table arranged? *Increasing atomic number*
- What are the vertical columns called? *Groups or Families*
- What are the horizontal rows called? *Periods*
- Why are elements put into groups? *Same number of valence electrons*
- What is a valence electron? *Number of electrons on the outer energy level (shell)*
- How many valence electrons each group has
- Which group of metals are highly reactive and why? *Alkali Metals. Only 1 val for loss*
- Which group of non-metals are highly reactive and why? *Halogens. Only 1 val to gain*
- Name some physical properties of metals: *luster, ductile, malleable, conductor of heat and electricity*
- Name some physical properties of non-metals: *no luster, brittle, not a good conductor*
- What are metalloids and where are they found on the periodic table? *Share properties of metals and metalloids. On a zigzag line starting at Boron*

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JA Banks



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Table of Contents Nova Hunting The Elements Questions And Answers

1. Understanding the eBook Nova Hunting The Elements Questions And Answers
 - The Rise of Digital Reading Nova Hunting The Elements Questions And Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Nova Hunting The Elements Questions And Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Nova Hunting The Elements Questions And Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Nova Hunting The Elements Questions And Answers
 - Personalized Recommendations
 - Nova Hunting The Elements Questions And Answers User Reviews and Ratings
 - Nova Hunting The Elements Questions And Answers and Bestseller Lists
5. Accessing Nova Hunting The Elements Questions And Answers Free and Paid eBooks
 - Nova Hunting The Elements Questions And Answers Public Domain eBooks
 - Nova Hunting The Elements Questions And Answers eBook Subscription Services
 - Nova Hunting The Elements Questions And Answers Budget-Friendly Options
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 - Nova Hunting The Elements Questions And Answers Compatibility with Devices
 - Nova Hunting The Elements Questions And Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Nova Hunting The Elements Questions And Answers
 - Highlighting and Note-Taking Nova Hunting The Elements Questions And Answers
 - Interactive Elements Nova Hunting The Elements Questions And Answers
8. Staying Engaged with Nova Hunting The Elements Questions And Answers

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Nova Hunting The Elements Questions And Answers
- 9. Balancing eBooks and Physical Books Nova Hunting The Elements Questions And Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Nova Hunting The Elements Questions And Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Nova Hunting The Elements Questions And Answers
 - Setting Reading Goals Nova Hunting The Elements Questions And Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Nova Hunting The Elements Questions And Answers
 - Fact-Checking eBook Content of Nova Hunting The Elements Questions And Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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