

temperature:

- a measure of the average kinetic energy of the particles in a sample of matter
- does not depend on the amount of matter in the sample
- symbol is T; unit is °C

heat:

- total amount of kinetic energy that flows because of a difference in temperature
- depends on mass of sample
- symbol is Q; unit is J (1 J = 4.18 cal)

Kinetic energy is energy in motion

Potential energy is stored

- Potential energy is hiding and cannot be measured
- Only changes in P.E. can be measured.

specific heat capacity:

- amount of energy required to raise the temperature of 1 gram of substance 1 °C
- symbol is c; unit is J/g°C

$$\begin{array}{ccccc}
 & \text{mass} & & & \\
 & \downarrow & & & \\
 \text{heat} \nearrow & Q = m \times C \times \Delta t & \nwarrow & \text{change in} & \\
 & \text{specific heat} & & \text{temperature} &
 \end{array}$$

When heat (Q) is absorbed by a system, part of it (C) goes into storage as potential energy and part of it is used to make the molecules move around more, raising the temperature (Δt).

Why does sand get hotter in the day and colder at night than the water?

Note Taking Guide Episode 13physicsfundamentals

David Kirk



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Table of Contents Note Taking Guide Episode 13physicsfundamentals

1. Understanding the eBook Note Taking Guide Episode 13physicsfundamentals
 - The Rise of Digital Reading Note Taking Guide Episode 13physicsfundamentals
 - Advantages of eBooks Over Traditional Books
2. Identifying Note Taking Guide Episode 13physicsfundamentals
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in a Note Taking Guide Episode 13physicsfundamentals
 - User-Friendly Interface
4. Exploring eBook Recommendations from Note Taking Guide Episode 13physicsfundamentals
 - Personalized Recommendations
 - Note Taking Guide Episode 13physicsfundamentals User Reviews and Ratings
 - Note Taking Guide Episode 13physicsfundamentals and Bestseller Lists

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 - Note Taking Guide Episode 13physicsfundamentals Public Domain eBooks
 - Note Taking Guide Episode 13physicsfundamentals eBook Subscription Services
 - Note Taking Guide Episode 13physicsfundamentals Budget-Friendly Options
6. Navigating Note Taking Guide Episode 13physicsfundamentals eBook Formats
 - ePub, PDF, MOBI, and More
 - Note Taking Guide Episode 13physicsfundamentals Compatibility with Devices
 - Note Taking Guide Episode 13physicsfundamentals Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Note Taking Guide Episode 13physicsfundamentals
 - Highlighting and Note-Taking Note Taking Guide Episode 13physicsfundamentals
 - Interactive Elements Note Taking Guide Episode 13physicsfundamentals
8. Staying Engaged with Note Taking Guide Episode 13physicsfundamentals
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Note Taking Guide Episode 13physicsfundamentals
9. Balancing eBooks and Physical Books Note Taking Guide Episode 13physicsfundamentals
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Note Taking Guide Episode 13physicsfundamentals
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Note Taking Guide Episode 13physicsfundamentals
 - Setting Reading Goals Note Taking Guide Episode 13physicsfundamentals
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Note Taking Guide Episode 13physicsfundamentals
 - Fact-Checking eBook Content of Note Taking Guide Episode 13physicsfundamentals
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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