

temperature:

- a measure of the average kinetic energy of the particles in a sample of matter
- does not depend on the amount of matter in the sample
- symbol is T; unit is °C

heat:

- total amount of kinetic energy that flows because of a difference in temperature
- depends on mass of sample
- symbol is Q; unit is J (1 J = 4.18 cal)

Kinetic energy is energy in motion

Potential energy is store

- Potential energy is hiding and cannot be measured
- Only changes in P.E. can be measured.

specific heat capacity:

- amount of energy required to raise the temperature of 1 gram of substance 1 °C
- symbol is c; unit is J/g°C

$$Q = m \times C \times \Delta t$$

heat →
mass →
specific heat →
change in temperature →

When heat (Q) is absorbed by a system, part of it (C) goes into storage as potential energy and part of it is used to make the molecules move around more, raising the temperature ( $\Delta t$ ).

Why does sand get hotter in the day and colder at night than the water?

# Note Taking Guide Episode 1302 Chemistry Key

**CL Gary**



## **Note Taking Guide Episode 1302 Chemistry Key:**

## Reviewing **Note Taking Guide Episode 1302 Chemistry Key**: Unlocking the Spellbinding Force of Linguistics

In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is truly astonishing. Within the pages of "**Note Taking Guide Episode 1302 Chemistry Key**," an enthralling opus penned by a very acclaimed wordsmith, readers set about an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve into the book's central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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