

Calculus-Based Thermodynamics (Practice Problems)

1. What will be the freezing-point and boiling-point of an aqueous solution containing 10.0 g of glucose, $C_6H_{12}O_6$, and 100 g of water? Equilibrium is at 1 atm. $K_f = 1.86^\circ C/m$.

$$\text{Molality} = \frac{\text{moles solute}}{\text{kg solvent}} = \frac{0.177 \text{ mol}}{0.100 \text{ kg}} = 1.77 \text{ mol/kg}$$

$$\Delta T_f = K_f \cdot m = 1.86^\circ C/m \cdot 1.77 \text{ mol/kg} = 3.30^\circ C$$

$\Delta T_f = 1.86^\circ C/m \cdot 1.77 \text{ mol/kg} = 3.30^\circ C$
 $T_f = 0^\circ C - 3.30^\circ C = -3.30^\circ C$
 $\Delta T_b = K_b \cdot m = 0.51^\circ C/m \cdot 1.77 \text{ mol/kg} = 0.90^\circ C$
 $T_b = 100^\circ C + 0.90^\circ C = 100.90^\circ C$

2. What will be the freezing-point and boiling-point of an aqueous solution containing 1.00 g of urea, CH_4N_2O , and 100 g of water? Equilibrium is at 1 atm. $K_f = 1.86^\circ C/m$, $K_b = 0.51^\circ C/m$.

$$\text{Molality} = \frac{\text{moles solute}}{\text{kg solvent}} = \frac{0.037 \text{ mol}}{0.100 \text{ kg}} = 0.37 \text{ mol/kg}$$

$$\Delta T_f = K_f \cdot m = 1.86^\circ C/m \cdot 0.37 \text{ mol/kg} = 0.69^\circ C$$

$\Delta T_f = 1.86^\circ C/m \cdot 0.37 \text{ mol/kg} = 0.69^\circ C$
 $T_f = 0^\circ C - 0.69^\circ C = -0.69^\circ C$
 $\Delta T_b = K_b \cdot m = 0.51^\circ C/m \cdot 0.37 \text{ mol/kg} = 0.19^\circ C$
 $T_b = 100^\circ C + 0.19^\circ C = 100.19^\circ C$

3. A 100 mL sample of 0.10 M $NaCl$ is cooled to $-5.0^\circ C$ at 1 atm. Assuming the freezing-point depression constant, $K_f = 1.86^\circ C/m$, find:

$$\Delta T_f = K_f \cdot m \quad 5.0^\circ C = 1.86^\circ C/m \cdot m \quad m = 2.69 \text{ mol/kg}$$

$$\Delta T_f = K_f \cdot m \quad 5.0^\circ C = 1.86^\circ C/m \cdot m$$

$$m = \frac{5.0^\circ C}{1.86^\circ C/m} = 2.69 \text{ mol/kg}$$

$$m = 2.69 \text{ mol/kg}$$

$$m = \frac{\text{moles solute}}{\text{kg solvent}} \quad 2.69 \text{ mol/kg} = \frac{\text{moles solute}}{0.100 \text{ kg}}$$

$$m = 2.69 \text{ mol/kg} = \frac{\text{moles solute}}{0.100 \text{ kg}}$$

$$\text{moles solute} = 0.269 \text{ mol}$$

$$\text{moles solute} = 0.269 \text{ mol} \quad \text{mass solute} = 0.269 \text{ mol} \cdot 58.44 \text{ g/mol} = 15.6 \text{ g}$$

4. A solution composed from 10 g of glucose, $C_6H_{12}O_6$, and 1.00 g of water is cooled to $-5.0^\circ C$. The freezing-point of this solution is $-5.0^\circ C$. The normal freezing-point of pure water is $0^\circ C$. Determine the molal mass of the compound. Assume $K_f = 1.86^\circ C/m$.

$$\Delta T_f = K_f \cdot m \quad 5.0^\circ C = 1.86^\circ C/m \cdot m \quad m = 2.69 \text{ mol/kg}$$

$$m = \frac{\text{moles solute}}{\text{kg solvent}}$$

$$2.69 \text{ mol/kg} = \frac{\text{moles solute}}{0.001 \text{ kg}}$$

$$m = 2.69 \text{ mol/kg}$$

$$m = \frac{\text{moles solute}}{\text{kg solvent}} \quad 2.69 \text{ mol/kg} = \frac{\text{moles solute}}{0.001 \text{ kg}}$$

$$2.69 \text{ mol/kg} = \frac{\text{moles solute}}{0.001 \text{ kg}}$$

$$\text{moles solute} = 0.00269 \text{ mol}$$

$$\text{moles solute} = 0.00269 \text{ mol} \quad \text{mass solute} = 0.00269 \text{ mol} \cdot M = 10 \text{ g}$$

$$\text{moles solute} = 0.00269 \text{ mol} \quad \text{mass solute} = 10 \text{ g} \quad M = 3717 \text{ g/mol}$$

Rosengarten Colligative Properties Homework Answers

Julia Schneider



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Table of Contents Rosengarten Colligative Properties Homework Answers

1. Understanding the eBook Rosengarten Colligative Properties Homework Answers
 - The Rise of Digital Reading Rosengarten Colligative Properties Homework Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Rosengarten Colligative Properties Homework Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Rosengarten Colligative Properties Homework Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Rosengarten Colligative Properties Homework Answers
 - Personalized Recommendations

- Rosengarten Colligative Properties Homework Answers User Reviews and Ratings
- Rosengarten Colligative Properties Homework Answers and Bestseller Lists
- 5. Accessing Rosengarten Colligative Properties Homework Answers Free and Paid eBooks
 - Rosengarten Colligative Properties Homework Answers Public Domain eBooks
 - Rosengarten Colligative Properties Homework Answers eBook Subscription Services
 - Rosengarten Colligative Properties Homework Answers Budget-Friendly Options
- 6. Navigating Rosengarten Colligative Properties Homework Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Rosengarten Colligative Properties Homework Answers Compatibility with Devices
 - Rosengarten Colligative Properties Homework Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Rosengarten Colligative Properties Homework Answers
 - Highlighting and Note-Taking Rosengarten Colligative Properties Homework Answers
 - Interactive Elements Rosengarten Colligative Properties Homework Answers
- 8. Staying Engaged with Rosengarten Colligative Properties Homework Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Rosengarten Colligative Properties Homework Answers
- 9. Balancing eBooks and Physical Books Rosengarten Colligative Properties Homework Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Rosengarten Colligative Properties Homework Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Rosengarten Colligative Properties Homework Answers
 - Setting Reading Goals Rosengarten Colligative Properties Homework Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Rosengarten Colligative Properties Homework Answers
 - Fact-Checking eBook Content of Rosengarten Colligative Properties Homework Answers

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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