

## Polyatomic Ions Worksheet Answer Key

### Polyatomic Practice

1. Name or write the formula for the following polyatomic ions

|              |                                      |                                       |              |
|--------------|--------------------------------------|---------------------------------------|--------------|
| sulfate      | <u><math>\text{SO}_4^{2-}</math></u> | <u><math>\text{CO}_3^{2-}</math></u>  | carbonate    |
| nitrite      | <u><math>\text{NO}_2^-</math></u>    | <u><math>\text{MnO}_4^{1-}</math></u> | manganate    |
| perphosphate | <u><math>\text{PO}_4^{3-}</math></u> | <u><math>\text{SO}_5^{2-}</math></u>  | persulfate   |
| hypoiodite   | <u><math>\text{IO}^-</math></u>      | <u><math>\text{BrO}_2^{1-}</math></u> | bromite      |
| chlorite     | <u><math>\text{ClO}_2^-</math></u>   | <u><math>\text{CO}_4^{2-}</math></u>  | percarbonate |
| phosphite    | <u><math>\text{PO}_3^{3-}</math></u> | <u><math>\text{PO}_5^{3-}</math></u>  | perphosphate |
| percarbonate | <u><math>\text{CO}_4^{2-}</math></u> | <u><math>\text{ClO}^{1-}</math></u>   | hypochlorite |
| bromate      | <u><math>\text{BrO}_3^-</math></u>   | <u><math>\text{IO}_2^{1-}</math></u>  | iodite       |
| hyposulfite  | <u><math>\text{SO}_2^{2-}</math></u> | <u><math>\text{PO}_4^{3-}</math></u>  | phosphate    |
| permanganate | <u><math>\text{MnO}_4^-</math></u>   | <u><math>\text{NO}_2^{1-}</math></u>  | nitrite      |
| carbonite    | <u><math>\text{CO}_2^{2-}</math></u> | <u><math>\text{SO}_4^{2-}</math></u>  | sulfate      |

2. Name or write the formula for the following Type I polyatomic ionic compounds

|                         |                                                                 |                                                                 |                         |
|-------------------------|-----------------------------------------------------------------|-----------------------------------------------------------------|-------------------------|
| beryllium hydroxide     | <u><math>\text{Be}(\text{OH})_2</math></u>                      | <u><math>\text{Ba}(\text{IO}_3)_2</math></u>                    | barium iodate           |
| sodium nitrite          | <u><math>\text{NaNO}_2</math></u>                               | <u><math>\text{Ga}(\text{CNO})_3</math></u>                     | gallium cyanate         |
| ammonium chloride       | <u><math>\text{NH}_4\text{Cl}</math></u>                        | <u><math>\text{Ag}_2\text{SO}_3</math></u>                      | silver sulfite          |
| calcium bisulfate       | <u><math>\text{Ca}(\text{HSO}_4)_2</math></u>                   | <u><math>\text{MgCO}</math></u>                                 | magnesium hypocarbonite |
| rubidium perchlorate    | <u><math>\text{RbClO}_4</math></u>                              | <u><math>\text{NH}_3\text{NO}_2</math></u>                      | ammonium nitrite        |
| strontium sulfite       | <u><math>\text{SrSO}_3</math></u>                               | <u><math>\text{Al}(\text{C}_2\text{H}_3\text{O}_2)_3</math></u> | aluminum acetate        |
| aluminum acetate        | <u><math>\text{Al}(\text{C}_2\text{H}_3\text{O}_2)_3</math></u> | <u><math>\text{SrSO}_5</math></u>                               | strontium persulfate    |
| ammonium nitrate        | <u><math>\text{NH}_4\text{NO}_3</math></u>                      | <u><math>\text{RbClO}_2</math></u>                              | rubidium chlorite       |
| magnesium hypocarbonite | <u><math>\text{MgCO}</math></u>                                 | <u><math>\text{Ca}(\text{HSO}_4)_2</math></u>                   | calcium bisulfate       |
| silver hyposulfite      | <u><math>\text{Ag}_2\text{SO}_2</math></u>                      | <u><math>\text{NH}_4\text{Cl}</math></u>                        | ammonium chloride       |

# Polyatomic Ions Worksheet Answer Key Unit 3

**Wolfgang Guggemos**

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located to the right of the author's name.

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