

LR Chem 105 - Worksheet on Molarity

Rules of operation: Work in pairs: one person should do the writing, and the other should do the thinking out loud. You should switch off between questions. You may consult with another pair to check answers or to resolve questions.

Molarity = moles solute / L of solution

$$3.0 \text{ M } \text{NH}_3 = \left(\frac{3.0 \text{ moles } \text{NH}_3}{1 \text{ L}} \right)$$

①. How many moles of NH_3 are contained in 75 mL of a 3.0 M solution of NH_3 ?

$$(75 \text{ mL}) \left(\frac{1 \text{ L}}{1000 \text{ mL}} \right) \left(\frac{3.0 \text{ mol}}{1 \text{ L}} \right) = .23 \text{ moles}$$

②. What volume of a 3.0 M solution of NH_3 should be taken to have 0.225 moles of NH_3 ?

$$.225 \text{ moles } \text{NH}_3 \left(\frac{1 \text{ L}}{3.0 \text{ moles } \text{NH}_3} \right) = .075 \text{ L } (= 75 \text{ mL})$$

③. How many grams of NaCl are contained in 50.0 mL of a 0.85 M NaCl solution?

You can do this in two steps if you like, or as a chain calculation. Maybe the "Mole Map" at the end of this handout will help.

$$50.0 \text{ mL} \left(\frac{1 \text{ L}}{1000 \text{ mL}} \right) \left(\frac{.85 \text{ mol NaCl}}{1 \text{ L}} \right) \left(\frac{58.5 \text{ g NaCl}}{1 \text{ mol NaCl}} \right) = 2.5 \text{ g NaCl}$$

Pogil Molarity Worksheet

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Solving Molarity Problems, 2016 Extend your understanding of molarity by solving some typical problems encountered in the high school chemistry classroom To foster your understanding of these problems you are asked to draw upon the quantitative reasoning skills you previously used

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