Name _______Period ______ Date _____

Physical Science - 2nd Semester - Final Exam Study Guide - (2012-2013)

1. Physical Science Basics

a. What tool(s) would you want to use to find the:

Mass of an object? triple beam balance

ii. Volume of an object? ruler or water displacement method

iii. Length of an object? Ruler

iv. Density of an object? graduated cylinder

- Basic SI Unit? 🗽
- Basic SI Unit? cm1 or ml

Basic SI Unit? meter

Basic SI Unit? ml

- b. What is the formula for calculating density? mass/volume
 - What is the density of a cube that is 15 g and has a volume of 3 cm³? 5g/cm³
 - ii. Compare the density of solid H₂O to liquid H₂O, and explain what observable effect this has.

solid H₂Ois less dense than liquid H₂O

2. Chemistry Basics

Element (def.) substance made up of one kind of atom

b. Atom (def.) basic building block of matter

Compound (def.) __ a molecule that contains atoms or molecules of two or more different elements.

d. Molecule (def.) a combination of two or more of the same or different atoms that are bonded together

3. Model of the Atom: Evolution of the Atomic Theory

Atoms are so small that many scientists throughout history have created models to describe them.

Why have atomic theories changed over time? Better technology came around.

Scientist	Year	Major Discovery/ Contribution	Name of atomic model
John Dalton	1802	Each element is made up of small, indivisible particles called atoms. Different elements have atoms of different masses.	Billiard Ball Model
J.J. Thomson	1897	Negatively charged electron	Plum Pudding Model
Niels Bohr	1913	Electrons move <u>fixed orbits</u> around the atom's nucleus in specific layers, or <u>electron shells</u> ; Atoms absorb or give off energy when the electrons move from one shell (<u>energy level</u>) to another	Bohr Model
Ernest Rutherford	1909	1909 (Based on idea that positive reflects positive) atom mostly empty space; electrons orbit randomly around positively charged nucleus	Planetary Model

4. Periodic Table

- The current periodic table is arranged by increasing atomic number.
- Reading the Periodic Table
 - An element's properties can be predicted from its location in the periodic table.
 - ii. The atomic number represents the number of protons.
 - The atomic mass represents the number of neutrons and protons.

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SA Adler

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