

Chapter 1: Limiting Reagents and Percent Yield (Pages 1-3)

Remember: Limiting reagent is the reactant that will completely react and leave another reactant

Excess: All reactants except the limiting reagent will be left over. Remember, the limiting reagent is the reactant that will completely react.

Remember: Limiting reagent will determine how much product is produced and byproducts will be the leftover reactant in a chemical reaction.



The ratio required for the composition for the chemical reaction.

2H ₂ (g)	100.00 percent
O ₂ (g)	50.00 percent
2H ₂ O(l)	100.00 percent
H ₂ O	100.00 percent

- (1) 100.00 g of each element will be used & react to form 100.00 g of product water. 100.00 g of hydrogen & 50.00 g of oxygen will be used. 100.00 g of water is produced.



- (2) 100.00 g H₂(g) is the limiting reagent. 49.50 g O₂(g) is the excess reagent.

- (3) 100.00 g of hydrogen will be completely consumed & 49.50 g of oxygen will be present.



- (4) 100.00 g of hydrogen is present in a 100.00 g of hydrogen & 49.50 g of oxygen. 100.00 g of hydrogen is present in a 100.00 g of hydrogen & 49.50 g of oxygen. 100.00 g of hydrogen is present in a 100.00 g of hydrogen & 49.50 g of oxygen.



Percent Yield And Limiting Reactants Quiz

Shasha Hu



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