

Single-Replacement Reactions

Use with
Section 9.2

A single-replacement chemical reaction is one in which one substance from a compound is replaced by another substance. A generic equation for such a reaction is as follows.



The reactivity of a substance depends on its ability to gain or lose electrons. It is possible to arrange the elements into a series based upon their reactivity. Such a list is called an activity series.

While there are many types of replacement reactions, we will concern ourselves with two different kinds. In one type, a more active metal replaces a less active metal from solution. Consider the reaction between zinc and copper(II) sulfate.



In this reaction, the more active zinc replaces the less active copper from solution. The reaction is evident because the blue color of the copper sulfate solution slowly turns colorless and a deposit of copper can be seen to form on the strip of zinc.

A second type of replacement reaction involves the replacement of hydrogen from acid by a metal. Consider the reaction between zinc and hydrochloric acid.



The zinc metal is active enough to replace the hydrogen from the acid. Bubbles of hydrogen gas can be seen rising to the surface, and the piece of zinc is consumed. On the other hand, if the less active metal, copper, is placed into a hydrochloric acid solution, no reaction will take place.

In this activity, you will use a few metals, their compounds, and dilute hydrochloric acid to show single-replacement reactions and construct an activity series.

Problem

What elements will replace other elements in single-replacement reactions? How can the results of these reactions be used to form an activity series?

Objectives

- **Classify** reactions as single-replacement chemical reactions.
- **Use numbers** to write balanced equations for single-replacement reactions.
- **Sequence** metals into an activity series.

Materials

zinc (1-cm × 3-cm strips) (3)
copper (1-cm × 3-cm strips) (2)
lead (1-cm × 3-cm strip)
sandpaper
0.2M lead (II) nitrate ($\text{Pb}(\text{NO}_3)_2$)
0.2M copper (II) sulfate (CuSO_4)

0.2M magnesium sulfate (MgSO_4)
0.2M silver nitrate (AgNO_3)
3M hydrochloric acid (HCl)
test tubes (6)
test-tube rack

Replacement Reactions Lab Answer Key

Padhraic Smyth

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

Replacement Reactions Lab Answer Key:

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