## CLASSIFYING CHEMICAL REACTIONS Name \_\_\_

Classify the following reactions as synthesis, decomposition, single replacement or double replacement.

- 2KCIO<sub>3</sub> → 2KCI + 3O<sub>2</sub>
- 2. HCl + NaOH → NaCl + H<sub>2</sub>O
- 3. Mg + 2HCl → MgCl<sub>2</sub> + H<sub>2</sub>
- 2H<sub>2</sub> + O<sub>2</sub> → 2H<sub>2</sub>O
- 2AI + 3NiBr₂ → 2AIBr₃ + 3Ni
- 4AI + 3O₂ → 2AI₂O₃
- 7. 2NaCl → 2Na + Cl<sub>2</sub>
- 8.  $CaCl_2 + F_2 \rightarrow CaF_2 + Cl_2$
- 9. AgNO<sub>3</sub> + KCl → AgCl + KNO<sub>3</sub>
- 10.  $N_2 + 3H_2 \rightarrow 2NH_3$
- 2H<sub>2</sub>O<sub>2</sub> → 2H<sub>2</sub>O + O<sub>2</sub>
- 12.  $(NH_4)_2SO_4 + Ba(NO_3)_2 \rightarrow BaSO_4 + 2NH_4NO_3$
- 13.  $Mgl_2 + Br_2 \rightarrow MgBr_2 + I_2$
- 14. SO<sub>3</sub> + H<sub>2</sub>O → H<sub>2</sub>SO<sub>4</sub>
- 15.  $6KCI + Zn_3(PO_4)_2 \rightarrow 3ZnCI_2 + 2K_3PO_4$

# **Physical Science If8767 Answers Binary Compounds**

**L** Towne

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