

Periodic Trends Worksheet

Use the periodic table and your knowledge of periodic trends to answer the following questions.

1. Which atom in each pair has the larger atomic radius?

a) Li or **K** b) **Ca** or Ni c) **Ga** or B d) O or **C** e) Cl or **Br**
f) Be or **Ba** g) **Si** or S h) Fe or **Au**

2. What is the periodic trend for atomic size from top to bottom in a group? from left to right in a period?

Top to bottom: increases due to adding energy levels. Left to Right: decreases

3. Why do atoms get smaller as you move left to right in a period?

due to increased nuclear charge and no added shielding.

4. Which element in each pair has a larger ionization energy?

a) Na or **O** b) **Be** or Ba c) **Ar** or F d) **Cu** or Ra e) I or **Ne**
f) K or **V** g) **Ca** or Fr h) W or **Se**

5. Explain the relationship between the relative size of an ion to its neutral atom and the charge on the ions.

Positive ions are smaller than the neutral atom it comes from, negative ions are larger than the neutral atom it comes from

6. Which particle has the larger radius in each atom/ion pair?

a) Na, Na⁺ b) S, S²⁻ c) I, I⁻ d) Al, Al³⁺

7. What is ionization energy? What is first ionization energy?

The amount of energy needed to remove an electron from an atom. First ionization energy is the energy required to remove a single valence electron from an atom.

8. What is the periodic trend for first ionization energy?

It increases going from left to right due to increased nuclear charge, and decreases going down a group due to increased distance from the nucleus and increased shielding.

9. Arrange the following groups of elements in order of increasing ionization energy.

a) Be, Mg, Sr b) Bi, Cs, Ba c) Na, Al, S
high, middle, low high, low, middle low, middle, high

10. Which element in each pair has a higher electronegativity value?

a) Cl, **F** b) C, **N** c) **Mg**, Ne d) **As**, Ca

11. What is the periodic trend for electronegativity? Increases going left to right, decreases going down.

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Table of Contents Periodic Trends Extension Questions Answers

1. Understanding the eBook Periodic Trends Extension Questions Answers
 - The Rise of Digital Reading Periodic Trends Extension Questions Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Trends Extension Questions Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Trends Extension Questions Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Trends Extension Questions Answers
 - Personalized Recommendations
 - Periodic Trends Extension Questions Answers User Reviews and Ratings
 - Periodic Trends Extension Questions Answers and Bestseller Lists

5. Accessing Periodic Trends Extension Questions Answers Free and Paid eBooks
 - Periodic Trends Extension Questions Answers Public Domain eBooks
 - Periodic Trends Extension Questions Answers eBook Subscription Services
 - Periodic Trends Extension Questions Answers Budget-Friendly Options
6. Navigating Periodic Trends Extension Questions Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Periodic Trends Extension Questions Answers Compatibility with Devices
 - Periodic Trends Extension Questions Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodic Trends Extension Questions Answers
 - Highlighting and Note-Taking Periodic Trends Extension Questions Answers
 - Interactive Elements Periodic Trends Extension Questions Answers
8. Staying Engaged with Periodic Trends Extension Questions Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Trends Extension Questions Answers
9. Balancing eBooks and Physical Books Periodic Trends Extension Questions Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Trends Extension Questions Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Periodic Trends Extension Questions Answers
 - Setting Reading Goals Periodic Trends Extension Questions Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Periodic Trends Extension Questions Answers
 - Fact-Checking eBook Content of Periodic Trends Extension Questions Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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