

NAME _____

DATE _____

GRAPHING PERIODIC TRENDS

The Periodic Table is arranged according to the Periodic Law. The Periodic Law states that when elements are arranged in order of increasing atomic number, their physical and chemical properties show a periodic pattern.

Define the following terms(starting on page 170):

Atomic radius-

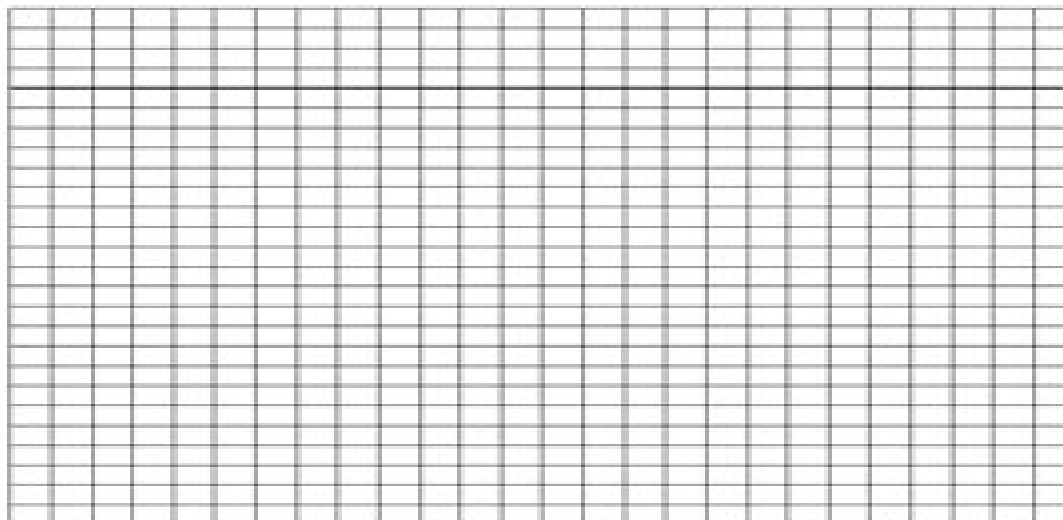
First ionization energy-

Electronegativity-

Graph 1 – Atomic Radius as a function of Atomic Number

- A. Create a graph of the atomic radius as a function of atomic number. Plot atomic number on the X axis and atomic radius on the Y axis. Remember to label the axes!
- B. Use a colored pen, pencil or highlighter to **trace** over the element's period (horizontal row on the periodic table). For example: use GREEN to **trace** for all of the elements in row 1, then use YELLOW to **trace** for all of the elements in row 2, then use ORANGE to **trace** for all the elements in row 3, then use BLUE to **trace** for all the elements in row 4.

Symbol	Atomic Number	Atomic Radius(pm)
H	1	37
He	2	50
Li	3	123
Be	4	89
B	5	80
C	6	77
N	7	70
O	8	66
F	9	64
Ne	10	67
Na	11	157
Mg	12	136
Al	13	125
Si	14	117
P	15	110
S	16	104
Cl	17	99
Ar	18	98
K	19	203
Ca	20	174



1. Describe the trend in atomic radius as the atomic number increases across a period.
2. Describe the trend in atomic radius as the atomic number increases down a group.

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C Cleary



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