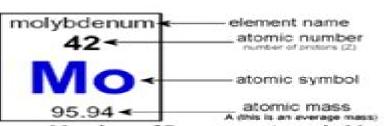
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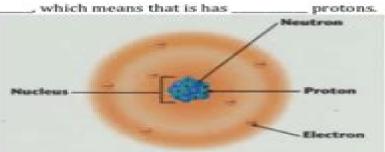
#### SPS1: Students will investigate our current understanding of the atom.

#### a. Examine the structure of the atom in terms of

- proton, electron, and neutron locations.
  - Proton has a charge of \_\_\_\_\_\_ and is located in the \_\_\_\_\_\_ of the atom.
    - Protons are also the elements number of the identity of the element, so protons are found by looking at the atomic number
  - Neutron has a charge of \_\_\_\_\_\_ and is located in the \_\_\_\_\_ of the atom.
    - Neutrons are found by subtracting the \_\_\_\_\_\_ Number from the \_\_\_\_\_\_ Mass or the BIG #-SMALL #.
  - Electron has a charge of \_\_\_\_\_ and is located in the \_\_\_\_ of the atom.
    - Electrons are equal to the number of \_\_\_\_\_\_ in a neutral atom.
  - The nucleus is made up of these 2 particles\_\_\_\_\_\_ and \_\_\_\_\_
  - The electrons are located in this region of the atom, known as the \_\_\_\_\_\_
  - The electron cloud is divided in to energy levels or energy shells, the first level will hold electrons and the second level will hold \_\_\_\_\_\_ electrons.
  - The valence shell or valence level is the \_\_\_\_\_\_ energy shell of the atom.
- atomic mass and atomic number.
  - atoms identity-it's the small number of the periodic table.
  - Atomic mass is equal to the average mass of all the isotopes of a particular element—it the big number on the periodic table.
- atoms with different numbers of neutrons (isotopes).
  - Define Isotope
  - Ex. Carbon-12 and Carbon-14
    - Carbon 12 has \_\_\_\_P \_\_\_N \_\_\_E Carbon 14 has \_\_\_P \_\_\_N \_\_\_E
- explain the relationship of the proton number to the element's identity.
  - The \_\_\_\_\_\_ number is equal to the number of \_\_\_\_\_ in the atom which is how the element is identified.
  - Ex. Nitrogen has an atom number of , which means that is has protons.



- Number of Protons = Atomic Number
- Number of Electrons = Number of Protons = Atomic Number
- Number of Neutrons = Mass Number Atomic Number



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**R Barnett** 

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