Name:

Date:

ROMEO & JULIET

Quotes Catalogue

Our coursework question will focus on Shakespeare's use of <u>opposing forces</u> throughout the play. He uses a wide range of oppositions to reflect the various forms of conflict and to explore a variety of themes (some of these can be further divided up into more opposing forces). The main opposing forces in the play are:



- Fate and Free Will
- Public and Private
- Appearance and Reality
- Loyalty and Betrayal
- NB You can find lots of information about these at the back of your books

Use this 'Quotes Catalogue' to make a note of the key quotations/scene numbers that reflect each of the opposing themes. Make sure you always have your quotes catalogue out in class and be ready to write down any key quotations that come up in discussion.

The more information you put into this catalogue, the easier your coursework will become!

Context

The big difference between this essay and your previous pieces of coursework is the inclusion of context. Assessment Objective 4 says you must:



This is how the exam board defines context. In other words, these are the things you should be thinking about when you write your essay:

- The historical setting, time and location of the text
- Social and cultural contexts (for example, attitudes in society; expectations of different cultural groups)
- The literary context of the text, for example, literary movements or genres
- The author's own life and individual situation, including the place and time of writing, only where these relate to the text

On your grid there is also a column to make a note of any contextual links you can make.

Romeo And Juliet Igcse Past Papers

M Lipman

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apr 25 2017 this is understood when you consider the two chemicals appear unchanged before and after the reaction therefore the net ionic equation is 2cl aq 2ag aq

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ions that remain essentially unchanged during a reaction are called spectator ions this means you can ignore them when you write the ionic equation you only need to model how the solid

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the first step to writing a net ionic equation is balancing the chemical equation present let's use the reaction between sodium chloride and silver nitrate as an example next we write the

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the net ionic equation is the chemical equation that shows only those elements compounds and ions that are directly involved in the chemical reaction notice that in writing the net ionic

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write the overall chemical equation the complete ionic equation and the net ionic equation for the reaction of aqueous silver fluoride with aqueous sodium phosphate to give solid silver

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the complete ionic equation for this reaction is as follows equation 4 13 2ag aq 2f aq 2nh4 aq cr2o72 aq ag2cr2o7 s 2nh4 aq 2f aq because two nh 4 aq

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