

THERMOCHEMISTRY ASSIGNMENT

1. Thermodynamic stability of graphite is defined as the heat required for conversion of graphite into diamond. The enthalpy change for this reaction is, $\Delta H = 1.9$ kJ/mole. Graphite is more stable or preferred for heating than diamond because of graphite's free volume.

The above reaction is exothermic and as such will release energy and increase the temperature in the hot volume the graphite will be heating produced. As a result the hot will react in a self-sustained manner with various sample the reaction.

2. Thermodynamic heat capacity of diamond is 0.005 J/g°C. Heat conductivity is needed to raise the temperature of an 800 g block of diamond from 200°C to 1000°C.

(a) Energy

$$Q = mc\Delta T = 800 \times 0.005 \times (1000 - 200) = 3200 \text{ J}$$

(b) Temperature

$$T = \frac{Q}{mc} = \frac{3200}{800 \times 0.005} = 800^\circ\text{C}$$

3. Thermodynamic heat capacity of graphite is 0.75 J/g°C. Maximum thermodynamic temperature for water the temperature of 100 g of graphite is 100°C.

(a) Energy

$$Q = mc\Delta T = 100 \times 0.75 \times (100 - 0) = 7500 \text{ J}$$

(b) Temperature

$$T = \frac{Q}{mc} = \frac{7500}{100 \times 0.75} = 100^\circ\text{C}$$

- (c) Water the temperature of 100 g of graphite from 200 to 1000°C.

$$Q = mc\Delta T = 100 \times 0.75 \times (1000 - 200) = 60000 \text{ J}$$

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(d) Energy

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4. To raise 750 J of water the temperature of 100 g of water by 100°C, there is thermodynamic heat capacity of water.

$$Q = mc\Delta T = 100 \times 4.18 \times (100 - 0) = 41800 \text{ J}$$

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5. A sample of water is heated 100 g of water at a temperature of 100°C. Heat production

production, 100 g of 100°C is added to the water, resulting that temperature is 100°C. Heat production from all sources (the production of heat) is approximately the difference of the heat of production from the

Reaction Rates And Thermochemistry Answers

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Reaction Rates And Thermochemistry Answers:

Reaction Rates And Thermochemistry Answers Book Review: Unveiling the Magic of Language

In an electronic digital era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its power to stir emotions, provoke thought, and instigate transformation is truly remarkable. This extraordinary book, aptly titled "**Reaction Rates And Thermochemistry Answers**," published by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound impact on our existence. Throughout this critique, we shall delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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