

Name: _____ per: _____

Worksheet - Reaction Rates

Use this reaction for the questions below: $C_6H_{12}O_6 (s) + 6 O_2(g) \rightarrow 6 H_2O (g) + 6 CO_2 (g)$

- What happens to the concentrations of:
 - $C_6H_{12}O_6$ & O_2 as the reaction proceeds \rightarrow ?
 - H_2O + CO_2 as the reaction proceeds \rightarrow ?
- According to the collision theory, what 3 circumstances are needed for $C_6H_{12}O_6$ & O_2 to react?
- What is the activation energy for a chemical reaction?
- Use the equation & the collision theory to explain:



Change in condition:	Does this <u>increase</u> or <u>decrease</u> the rate of reaction?	Explain why
a. <u>Increasing the temperature</u>	Ex: Increase (springs up)	Ex: Molecules move faster & collide more = increased rxn rate
b. <u>Increasing the concentration</u> of $C_6H_{12}O_6$		
c. <u>Decreasing the concentration</u> of O_2		
d. <u>Increase the surface area</u> by chewing up food in your mouth.		
e. <u>Decreasing the temperature</u>		
f. <u>Increasing the pressure</u> in the container		
g. <u>Decreasing the concentration</u> of H_2O		
h. <u>Increasing the volume</u> of the container the reaction occurs in		
i. <u>Increasing the concentration</u> of CO_2		
j. <u>Using a catalyst</u> (like salivary amylase)		

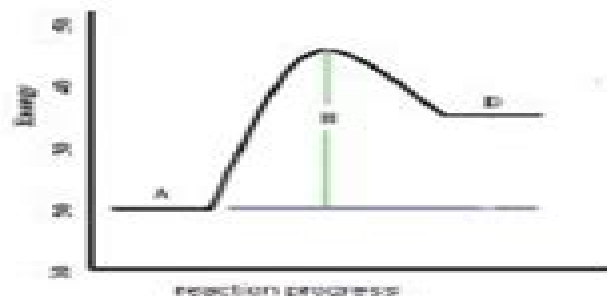
- On the accompanying energy diagram, label the following terms:

a. reactants b. products c. activation energy

- On the graph to the right, draw and label what this diagram would look like if a catalyst was added to the reaction.

- Graph reading

- How much energy (kJ) do the reactants have?
- How much energy (kJ) do the products have?
- How much energy (kJ) is required to activate this un-catalyzed reaction?
- Is this reaction endothermic or exothermic? How do you know?



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Holt Chemistry Holt Rinehart & Winston, 2003-01-24 **Science Spectrum** Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2003-03 Books in Print Supplement, 2002 *Ate Science Plus 2002 LV Red* Holt Rinehart & Winston, 2001-02 *Workbook 15* Thobeka Pete, Turret Correspondence College (Johannesburg), 1988 **Rates of Reaction & Equilibrium** Roy Brown, Irish science teachers' association, 1983 **Rates and Equilibria of Organic Reactions as Treated by Statistical, Thermodynamic, and Extrathermodynamic Methods** John E. Leffler, Ernest Grunwald, 1989-01-01 Graduate level text stresses extrathermodynamic approach to quantitative prediction and constructs a logical framework that encompasses and classifies all known extrathermodynamic relationships Numerous figures and tables Author and Subject Indexes *Rates and Mechanisms of Chemical Reactions* W. C. Gardiner (Jr.), William Cecil Gardiner, 1969 **Guide to the Rates of Reactions & Chemical Equilibrium Kit** University of Natal. Centre for the Advancement of Science and Mathematics Education, Shell South Africa (Pty) Limited, 1994 **Theories of Chemical Reaction Rates** Keith James Laidler, 1969 **An Introduction to Chemical Equilibrium and Kinetics** Louis Meites, 1981

Rates of Reaction and Equilibria Clifford Othen, 1972-01-01 Reaction Kinetics Keith J. Laidler, 2013-10-22

Reactions Kinetics Volume I Homogeneous Gas Reactions presents a general introduction to the subject of kinetics including the basic laws of kinetics and the theoretical treatment of reaction rates This four chapter book deals mainly with homogeneous reactions in the gas phase Chapter 1 presents the kinetic laws based on experimental results in terms of their simple concepts with a special consideration of the way in which rates depend on concentration while Chapter 2 deals with the interpretation of rates in terms of more fundamental theories Chapter 3 covers the overall reactions that are believed to be elementary such as the reaction between hydrogen and iodine the reverse decomposition of hydrogen iodide the corresponding reactions involving deuterium instead of hydrogen and the dimerizations of butadiene and cyclopentadiene as well as a few elementary termolecular reactions all involving nitric oxide This chapter also includes a general account of some of the elementary reactions that occur as steps in more complex mechanisms Chapter 4 examines the reaction rates of numerous complex gas reactions Undergraduate physical chemistry and chemical kinetics students as well as advanced students in other fields such as biology and physics will find this book invaluable **An Introduction to Chemical Kinetics** Claire Vallance, 2017-09-28 The book is a short primer on chemical reaction rates based on a six lecture first year undergraduate course taught by the author at the University of Oxford The book explores the various factors that determine how fast or slowly a chemical reaction proceeds and describes a variety of experimental methods for measuring reaction rates The link between the reaction rate and the sequence of steps that makes up the reaction mechanism is also investigated Chemical reaction rates is a core topic in all undergraduate chemistry courses Holt Science Spectrum Physical Science Chapter 7 Resource File: Chemical Reactions Holt Rinehart & Winston, 2008-01-01 *Chemical Reactions Science Learning*

Guide NewPath Learning,2014-03-01 The Chemical Reactions Student Learning Guide includes self directed readings easy to follow illustrated explanations guiding questions inquiry based activities a lab investigation key vocabulary review and assessment review questions along with a post test It covers the following standards aligned concepts Changes of Matter Chemical Reactions Formulas Balancing Equations Types of Chemical Reactions 1 Types of Chemical Reactions 2 Energy in Chemical Reactions Evidence of Chemical Reactions and Chemical Reaction Rates Catalysts Aligned to Next Generation Science Standards NGSS and other state standards Rate of Reaction, Sensitivity and Chemical Equilibrium American Institute of Chemical Engineers,1980-12-01 **The Equilibrium Assumption in the Theory of Absolute Reaction Rates** Charles Francis Curtiss,1948 The assumption that systems in the activated state are in equilibrium with the reactant molecules has been open to question In this paper the methods of classical mechanics are used to obtain an expression for the rate constant without making this assumption The result is similar in form to that obtained by Eyring except that the average velocity of the points in configuration space is related directly to the distribution in velocities of the molecules of the gas in real space If one assumes that this distribution is Maxwellian one obtains the Eyring expression However the chemical reaction itself acts as a perturbing influence and causes the distribution in velocities to be somewhat different from Maxwellian It is this perturbation of the distribution function which leads to a different value of the rate constant and consequently to a measure of the effect of the equilibrium assumption Author How Chemical Reactions Occur Edward L. King,1963 **Chemical Equilibrium** William Guenther,2012-12-06 The present work is designed to provide a practical introduction to aqueous equilibrium phenomena for both students and research workers in chemistry biochemistry geochemistry and interdisciplinary environmental fields The pedagogical strategy I have adopted makes heavy use of detailed examples of problem solving from real cases arising both in laboratory research and in the study of systems occurring in nature The procedure starts with mathematically complete equations that will provide valid solutions of equilibrium problems instead of the traditional approach through approximate concentrations and idealized infinite dilution assumptions There is repeated emphasis on the use of corrected conditional equilibrium constants and on the checking of numerical results by substitution in complete equations and or against graphs of species distributions Graphical methods of calculation and display are used extensively because of their value in clarifying equilibria and in leading one quickly to valid numerical approximations The coverage of solution equilibrium phenomena is not however exhaustively comprehensive Rather I have chosen to offer fundamental and rigorous examinations of homogeneous step equilibria and their interactions with solubility and redox equilibria Many examples are worked out in detail to demonstrate the use of equilibrium calculations and diagrams in various fields of investigation

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