Try to check how much you have learned in this lesson by answering the following questions.

- Write the equilibrium expression for the following reactions.
 - Ni_M + 400_{pr} ++ Ni(00)_{min}
 - 2. H₂O₁₀ ++ H₂O₁₀
 - 3. NOW + COM ++ COM + Now
 - 4. Haist + lass ++ 2HI_{bit}
 - 5. 3Hzur + Nzur ++ 2NHzur
- II. Consider the exothermic reaction

Write -+ if the reaction will go forward, +- for reverse and ++ for no change in the direction of the reaction.

- Adding oxygen
 Removing suffur dioxide
 Removing suffur broxide
- Adding helium gas
 Adding catalyst



Key to answers on page 15.

Lesson 3. Applications of Chemical Equilibrium

In this lesson, we will study some of the wonderful applications of chemical equilibrium. The principles of chemical equilibrium are evident from the production of industrial chemicals, the acid-balance in takes and streams, to the maintenance of our blood pH.

Haber Process: Ammonia synthesis.

We have encountered the synthesis of ammonia from previous lessons. This time, we will take a closer lock at this reaction. The complete equation is written as

Reaction Rates And Equilibrium Lesson Check Answers

Daniel C. Tofan

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Book of Chemistry Teacher Stories Jeff Lark, Stories from years of teaching high school chemistry *General Chemistry Workbook* Daniel C. Tofan, 2010-07-28 This workbook is a comprehensive collection of solved exercises and problems typical to AP introductory and general chemistry courses as well as blank worksheets containing further practice problems and questions It contains a total of 197 learning objectives grouped in 28 lessons and covering the vast majority of the types of

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