

## Equilibrium Practice Problems

1. Write the equilibrium expression for each of the following reactions:



$$K = \frac{[\text{NH}_3]^2}{[\text{N}_2] [\text{H}_2]^3}$$



$$K = \frac{[\text{ICl}]^2}{[\text{Cl}_2]}$$



$$K = \frac{[\text{NO}] [\text{O}_2]^{1/2}}{[\text{NO}_2]}$$

2. The dissociation of acetic acid,  $\text{CH}_3\text{COOH}$ , has an equilibrium constant at  $25^\circ\text{C}$  of  $1.8 \times 10^{-5}$ . The reaction is



If the equilibrium concentration of  $\text{CH}_3\text{COOH}$  is 0.46 moles in 0.500 L of water and that of  $\text{CH}_3\text{COO}^-$  is  $8.1 \times 10^{-3}$  moles in the same 0.500 L, calculate  $[\text{H}^+]$  for the reaction.

$$K = \frac{[\text{CH}_3\text{COO}^-] [\text{H}^+]}{[\text{CH}_3\text{COOH}]} \quad \text{Plug in known values and solve.}$$

$$1.8 \times 10^{-5} = \frac{[8.1 \times 10^{-3} \text{ moles}/0.500 \text{ L}] [\text{H}^+]}{[0.46 \text{ moles}/0.500 \text{ L}]}$$

$$[\text{H}^+] = 1.0 \times 10^{-3} \text{ M}$$

3. Indicate the effect of a catalyst, pressure, temperature and concentration on each of the following on:

	catalyst	pressure	temperature	concentration
a. speed	presence increase	direct relationship	direct relationship	direct relationship
b. pos <sup>n</sup>	nothing	based on moles	based on enthalpy of rxn	inverse relationship

# Reaction Rates And Equilibrium Practice Problems Answers

**G. F. Azzone**



## Reaction Rates And Equilibrium Practice Problems Answers:

**Survival Guide to General Chemistry** Patrick E. McMahon, Rosemary McMahon, Bohdan Khomtchouk, 2019-02-13 This work evolved over thirty combined years of teaching general chemistry to a variety of student demographics. The focus is not to recap or review the theoretical concepts well described in the available texts. Instead, the topics and descriptions in this book make available specific detailed step by step methods and procedures for solving the major types of problems in general chemistry. Explanations, instructional process sequences, solved examples, and completely solved practice problems are greatly expanded, containing significantly more detail than can usually be devoted to in a comprehensive text. Many chapters also provide alternative viewpoints as an aid to understanding. Key Features: The authors have included every major topic in the first semester of general chemistry and most major topics from the second semester. Each is written in a specific and detailed step by step process for problem solving, whether mathematical or conceptual. Each topic has greatly expanded examples and solved practice problems containing significantly more detail than found in comprehensive texts. Includes a chapter designed to eliminate confusion concerning acid base reactions, which often persists through working with acid base equilibrium. Many chapters provide alternative viewpoints as an aid to understanding. This book addresses a very real need for a large number of incoming freshmen in STEM fields.

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*Partial Solutions Manual* Darrell D. Ebbing, David Bookin, 1998

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*Lewin's CELLS* Lynne Cassimeris, Vishwanath Lingappa, George Plopper, Benjamin Lewin, 2011-03-25 Completely revised and updated to incorporate the latest data in the field Lewin's *CELLS* Second Edition is the ideal resource for advanced undergraduate and graduate students entering the world of cell biology Redesigned to incorporate new learning tools and elements this edition continues to provide readers with current coverage of the structure organization growth regulation movements and interaction of cells with an emphasis on eukaryotic cells Under the direction of three expert lead editors new chapters on metabolism and general molecular biology have been added by subject specialist All chapters have been carefully edited to maintain consistent use of terminology and to achieve a homogenous level of detail and rigor A new design incorporates many new pedagogical elements including Concept Reasoning Questions Methods boxes Clinical Applications boxes and more

*CliffsNotes AP Chemistry* Bobrow Test Preparation Services, 2009-02-09 The book itself contains chapter length subject reviews on every subject tested on the AP Chemistry exam as well as both sample multiple choice and free response questions at each chapter's end Two full length practice tests with detailed answer explanations are included in the book

*Excel With Objective Questions In Chemistry* Prof. S. K. Khanna, Dr. N. K. Verma, Dr. B. Kapila, 2006

*Microwaves in Organic Synthesis* Antonio de la Hoz, André Loupy, 2013-02-26 The third edition of the bestselling two volume reference covers everything you need to know about microwave technology for synthesis from the best equipment to nonthermal effects from solid support reactions to catalysis Completely revised and updated with half of the authors completely new to the project this comprehensive work is clearly divided into two parts on the fundamentals of microwave irradiation and application of microwaves and synergies with other enabling techniques Also new to this edition are chapters on on line monitoring flow chemistry combination with ultrasounds and natural products including multicomponent reactions An indispensable source for organic catalytic physical and medicinal chemists

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Colligative Properties Kinetics Equilibrium Electrochemistry Nuclear Chemistry Organic Chemistry Experimental **IIT**  
**Chemistry-II**, **Ebbing** Darrell D. Ebbing, 1993



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