



## Reaction Rates and Equilibrium Study Guide

### Topics Covered:

Factors affecting reaction rates  
Collision theory of reaction rates  
Reaction progress diagrams (activation energy and transition states)  
Rate laws and the method of initial rates  
Chemical equilibrium  
Equilibrium expressions and constants and their meaning  
Product-favored vs. Reactant-favored  
Shifting equilibrium and Le Chatelier's Principle  
Quantitative calculations of equilibrium systems using ICE tables

### Practice Questions: [Click here for an answer key](#)

- Describe how each of the following factors would affect the rate of "particle collisions" in your response.
  - Temperature decrease
  - Addition of iodide ions,  $I^-$ , as a catalyst
  - Increased reactant concentration
  - Crushing a solid into a powder (with larger surface area)
- The reaction progress diagram shows the energy changes that take place in the reaction of carbon monoxide gas with nitrogen dioxide gas.



# Reaction Rates And Equilibrium Study Guide

**Tao Wei**



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