Worksheet- Reaction Rates

Use this reaction for the questions below: $C_0H_{22}O_0$ (s) + 6 $O_2(g) \rightarrow$ 6 H_2O (g) + 6 CO_2 (g)

- 1. What happens to the concentrations of:
 - a. C₆H₁₂O₆ & O₂ as the reaction proceeds →?
- b. H₂O + CO₂ as the reaction proceeds →?
- 2. According to the collision theory, what 3 circumstances are needed for CoHizOo & Oz to react?
- 3. What is the activation energy for a chemical reaction?
- 4. Use the equation & the collision theory to explain:

$$C_6H_{12}O_6(s) + 6 O_2(g) \rightarrow 6 H_2O(g) + 6 CO_2(g)$$

	Change in condition:	Does this <u>increase</u> or <u>decrease</u> the rate of reaction?	Explain why
a.	Increasing the temperature	Exi (normanca (apenda up)	Ext Molecules more faster to collide more - Increased on rate
ь.	Increasing the concentration of CoHorOs		
C.	Decreasing the concentration of O ₂		
d.	Increase the surface area by chewing up food in your mouth		
6.	Decreasing the temperature		
f.	Increasing the pressure in the container		
9	Decreasing the concentration of H ₂ O		
h.	Increasing the volume of the container the reaction occurs in		
	Increasing the concentration of CO ₂		
1	Using a catalyst (like solivary amylase)		

- On the accompanying energy diagram, <u>label</u> the following terms:
 - a. reactants b. products c. activation energy
- On the graph to the right, <u>draw</u> and <u>label</u> what this diagram would look like if a <u>catalyst</u> was added to the reaction.
- 7. Graph reading
 - a. How much energy (#) do the reactants have?
 - b. How much energy (#) do the products have?
 - c. How much energy (#) is required to activate this un-catalyzed reaction?
 - d. Is this reaction endothermic or exothermic? How do you know?



Reaction Rates Equilibrium Worksheet Answers Chapter 19

R Bogdan

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Table of Contents Reaction Rates Equilibrium Worksheet Answers Chapter 19

- 1. Understanding the eBook Reaction Rates Equilibrium Worksheet Answers Chapter 19
 - The Rise of Digital Reading Reaction Rates Equilibrium Worksheet Answers Chapter 19
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Reaction Rates Equilibrium Worksheet Answers Chapter 19
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Reaction Rates Equilibrium Worksheet Answers Chapter 19
 - User-Friendly Interface

- 4. Exploring eBook Recommendations from Reaction Rates Equilibrium Worksheet Answers Chapter 19
 - Personalized Recommendations
 - Reaction Rates Equilibrium Worksheet Answers Chapter 19 User Reviews and Ratings
 - Reaction Rates Equilibrium Worksheet Answers Chapter 19 and Bestseller Lists
- 5. Accessing Reaction Rates Equilibrium Worksheet Answers Chapter 19 Free and Paid eBooks
 - Reaction Rates Equilibrium Worksheet Answers Chapter 19 Public Domain eBooks
 - Reaction Rates Equilibrium Worksheet Answers Chapter 19 eBook Subscription Services
 - Reaction Rates Equilibrium Worksheet Answers Chapter 19 Budget-Friendly Options
- 6. Navigating Reaction Rates Equilibrium Worksheet Answers Chapter 19 eBook Formats
 - o ePub, PDF, MOBI, and More
 - Reaction Rates Equilibrium Worksheet Answers Chapter 19 Compatibility with Devices
 - Reaction Rates Equilibrium Worksheet Answers Chapter 19 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - o Adjustable Fonts and Text Sizes of Reaction Rates Equilibrium Worksheet Answers Chapter 19
 - Highlighting and Note-Taking Reaction Rates Equilibrium Worksheet Answers Chapter 19
 - o Interactive Elements Reaction Rates Equilibrium Worksheet Answers Chapter 19
- 8. Staying Engaged with Reaction Rates Equilibrium Worksheet Answers Chapter 19
 - o Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Reaction Rates Equilibrium Worksheet Answers Chapter 19
- 9. Balancing eBooks and Physical Books Reaction Rates Equilibrium Worksheet Answers Chapter 19
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Reaction Rates Equilibrium Worksheet Answers Chapter 19
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Reaction Rates Equilibrium Worksheet Answers Chapter 19
 - Setting Reading Goals Reaction Rates Equilibrium Worksheet Answers Chapter 19
 - Carving Out Dedicated Reading Time

- 12. Sourcing Reliable Information of Reaction Rates Equilibrium Worksheet Answers Chapter 19
 - Fact-Checking eBook Content of Reaction Rates Equilibrium Worksheet Answers Chapter 19
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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