



Rds Eon Radio Manual

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apr 25 2017 this is understood when you consider the two chemicals appear unchanged before and after the reaction therefore the net ionic equation is $2\text{Cl}^- \text{aq} + 2\text{Ag}^+ \text{aq}$

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ions that remain essentially unchanged during a reaction are called spectator ions this means you can ignore them when you write the ionic equation you only need to model how the solid

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1 $\text{K}^+ \text{PO}_4^{3-} \text{aq} + \text{Al}^{3+} \text{aq} + \text{PO}_4^{3-} \text{s} + 3 \text{KNO}_3 \text{aq} + 3 \text{H}^+ \text{aq} + 3 \text{NO}_3^- \text{aq} + 3 \text{H}_2\text{O} \text{l}$ net ionic $\text{Al}^{3+} \text{aq} + 3 \text{PO}_4^{3-} \text{aq} \rightarrow \text{AlPO}_4 \text{s}$

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the complete ionic equation for this reaction is as follows equation 4 13 2ag aq 2f aq 2nh₄ aq cr₂o₇²⁻ aq ag₂cr₂o₇ s 2nh₄ aq

2f aq because two nh 4 aq
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