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# FORMULAS WITH POLYTOMIC IONS

Matching the horizontal and vertical axes, write the formulas of the compounds with the following combination of ions.

|                  | $\text{OH}^-$                          | $\text{OH}_2^-$            | $\text{CO}_3^{2-}$           | $\text{SO}_4^{2-}$           | $\text{PO}_4^{3-}$           |
|------------------|--|----------------------------|------------------------------|------------------------------|------------------------------|
| $\text{H}^+$     | $\text{HOH}$<br>$(\text{H}_2\text{O})$ | $\text{HNO}_3$             | $\text{H}_2\text{CO}_3$      | $\text{H}_2\text{SO}_4$      | $\text{H}_3\text{PO}_4$      |
| $\text{Na}^+$    | $\text{NaOH}$                          | $\text{NaNO}_3$            | $\text{Na}_2\text{CO}_3$     | $\text{Na}_2\text{SO}_4$     | $\text{Na}_3\text{PO}_4$     |
| $\text{Mg}^{2+}$ | $\text{Mg}(\text{OH})_2$               | $\text{Mg}(\text{NO}_3)_2$ | $\text{MgCO}_3$              | $\text{MgSO}_4$              | $\text{Mg}_3(\text{PO}_4)_2$ |
| $\text{NH}_4^+$  | $\text{NH}_4\text{OH}$                 | $\text{NH}_4\text{NO}_3$   | $(\text{NH}_4)_2\text{CO}_3$ | $(\text{NH}_4)_2\text{SO}_4$ | $(\text{NH}_4)_3\text{PO}_4$ |
| $\text{Ca}^{2+}$ | $\text{KOH}$                           | $\text{KNO}_3$             | $\text{K}_2\text{CO}_3$      | $\text{K}_2\text{SO}_4$      | $\text{K}_3\text{PO}_4$      |
| $\text{K}^+$     | $\text{Al}(\text{OH})_3$               | $\text{Al}(\text{NO}_3)_3$ | $\text{Al}_2(\text{CO}_3)_3$ | $\text{Al}_2(\text{SO}_4)_3$ | $\text{AlPO}_4$              |
| $\text{Al}^{3+}$ | $\text{Pb}(\text{OH})_4$               | $\text{Pb}(\text{NO}_3)_4$ | $\text{Pb}(\text{CO}_3)_2$   | $\text{Pb}(\text{SO}_4)_2$   | $\text{Pb}_3(\text{PO}_4)_4$ |

# Polyatomic Ionic Formulas Answers

**Lauren Gardner**



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