# Stoichiometry Problem Set A

Directions: Solve each of the following problems. Show work using the factor label method or the irf box. Include correct units and significant figures with each answer. Clearly indicate final answer. Be sure to start with a balanced equation

#### Part 1: Straight Stoichiometry Problems

Target Goal: To properly calculate the amount of one substance equal to a given amount of another substance within a chemical reaction using the factor-label method.

For the reaction HNO<sub>Neq</sub> + Mg(OH)<sub>100</sub> → Mg(NO<sub>2</sub>)<sub>3(eq)</sub> + H<sub>2</sub>O<sub>(0)</sub>, how many grams of magnesium nitrate are produced from 8.00 mol of nitric acid (HNO<sub>2</sub>)? A

MM=148.39

For the reaction 2 KClO<sub>300</sub> → 2 KCl<sub>00</sub> + 3 O<sub>200</sub>, how many moles of potassium chlorate are required to produce 160. g of oxygen? A

MH = 32.09

- KCID3
- 3.) For the reaction  $2H_{200} + O_{200} \rightarrow 2H_2O_{50}$ , how many grams of water are produces from 5.22 liters of oxygen gas?

MH= 18.09

4.) In the reaction Na<sub>(0)</sub> + PH<sub>2</sub>O<sub>(0)</sub> → NaOH<sub>0-0</sub> + H<sub>2(0)</sub>, how many grams of hydrogen are produced if 1.33 grams of sodium react completely?

MH= 23.09

5.) In the reaction \_\_\_\_CO<sub>2(a)</sub> + \_\_\_H<sub>2</sub>O<sub>00</sub> → \_\_\_H<sub>2</sub>CO<sub>2(a)</sub>, how many moles of carbonic acid can be produced from 97.2 g of CO<sub>2</sub>?

MH 44.09

2.21 moles Heca

#### Did you achieve our target goal?

Target Goal: To properly calculate the amount of one substance equal to a given amount of another substance within a chemical reaction using the factor-label method. Circle your answer.

Feeling VERY confused

3 I think I am catching on

Doing great! Bring it on!

# **Sample Problem Set Stoichometry Chemistry**

Neil D. Jespersen, Pamela Kerrigan

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#### **Table of Contents Sample Problem Set Stoichometry Chemistry**

- 1. Understanding the eBook Sample Problem Set Stoichometry Chemistry
  - The Rise of Digital Reading Sample Problem Set Stoichometry Chemistry
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Sample Problem Set Stoichometry Chemistry
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Sample Problem Set Stoichometry Chemistry
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Sample Problem Set Stoichometry Chemistry
  - Personalized Recommendations
  - Sample Problem Set Stoichometry Chemistry User Reviews and Ratings
  - Sample Problem Set Stoichometry Chemistry and Bestseller Lists
- 5. Accessing Sample Problem Set Stoichometry Chemistry Free and Paid eBooks

- Sample Problem Set Stoichometry Chemistry Public Domain eBooks
- Sample Problem Set Stoichometry Chemistry eBook Subscription Services
- Sample Problem Set Stoichometry Chemistry Budget-Friendly Options
- 6. Navigating Sample Problem Set Stoichometry Chemistry eBook Formats
  - o ePub, PDF, MOBI, and More
  - Sample Problem Set Stoichometry Chemistry Compatibility with Devices
  - Sample Problem Set Stoichometry Chemistry Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Sample Problem Set Stoichometry Chemistry
  - Highlighting and Note-Taking Sample Problem Set Stoichometry Chemistry
  - Interactive Elements Sample Problem Set Stoichometry Chemistry
- 8. Staying Engaged with Sample Problem Set Stoichometry Chemistry
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Sample Problem Set Stoichometry Chemistry
- 9. Balancing eBooks and Physical Books Sample Problem Set Stoichometry Chemistry
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Sample Problem Set Stoichometry Chemistry
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Sample Problem Set Stoichometry Chemistry
  - Setting Reading Goals Sample Problem Set Stoichometry Chemistry
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Sample Problem Set Stoichometry Chemistry
  - Fact-Checking eBook Content of Sample Problem Set Stoichometry Chemistry
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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