

DESCRIBING CHEMICAL REACTIONS

Section Review - Objectives

- Explain how to write a word equation
- Describe how to write a skeleton equation
- List the steps for writing a complete chemical equation.

Vocabulary

- chemical equation
- skeleton equation
- catalyst
- coefficients
- balanced equation

Part A Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short phrase, or number.

A chemical reaction can be concisely represented by a chemical ____ 1 _____. The substances that undergo a chemical change are the ____ 2 _____. The new substances formed in a chemical reaction are the ____ 3 _____. In accordance with the law of conservation of ____ 4 _____, a chemical equation must be balanced. When balancing an equation, you place ____ 5 _____ in front of reactants and products so that the same number of atoms of each ____ 6 _____ are on each side of the equation. An equation must never be balanced by changing the ____ 7 _____ in the chemical formula of a substance.

Special symbols are used to show the physical state of a substance in a reaction. The symbol for a liquid is ____ 8 ____; for a solid, ____ 9 ____; and for a gas, ____ 10 _____. A substance dissolved in water is designated ____ 11 _____. If a ____ 12 _____ is used to increase the rate of a chemical reaction, its formula is written above the arrow.

- _____
- _____
- _____
- _____
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- _____
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- _____
- _____
- _____

Part B True-False

Classify each of these statements as always true, AT; sometimes true, ST; or never true, NT.

- _____ 13. In an equation, a substance is shown to be in the gaseous state by placing an upward-pointing arrow after its formula.
- _____ 14. The symbol Δ placed over the arrow in an equation means that heat is supplied to the reaction.
- _____ 15. Atoms are destroyed in a chemical reaction.
- _____ 16. A skeleton equation is not a balanced equation.

Part D Questions and Problems

24. Write a balanced equation for each of these chemical reactions. I

a. Aluminum reacts with aqueous hydrochloric acid to form hydrogen gas and aqueous aluminum chloride.

b. Acetylene gas (C_2H_2) burns in a welding torch with oxygen to form carbon dioxide gas and water vapor.

Section 1describing Chemical Reactions Answer Key

RJ Shavelson



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