

Chapter 10 – CHEMICAL QUANTITIES

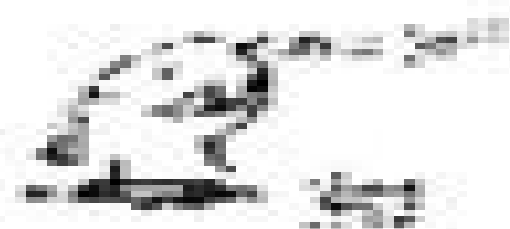
The MOLE

Avogadro's Hypothesis

Equal volumes of gases at constant T and p contain the same # molecules.

The number of H_2 atoms in 12.01 grams of carbon is called **Avogadro's Number** – 6.02×10^{23}

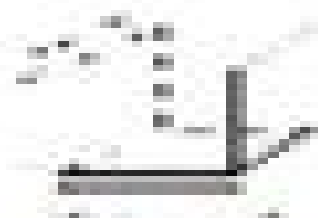
This quantity is called a **MOLE**.
(Just use a concept – IT)



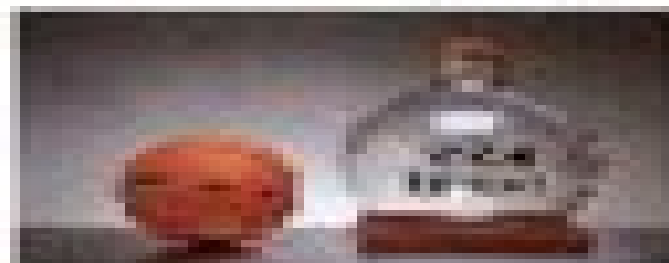
1 mole = 6.02×10^{23} particles (approximate conversion)

STP: Standard Temperature and Pressure
0°C 1 atm

At STP: **1 mole = 22.4 L** for any gas.



At STP, 1 mole is **22.4 L**,
equivalently 22.4 dm³ or 22.4 L



Section 10 Chemical Quantities Packet Answers

Rachel Sandford



Section 10 Chemical Quantities Packet Answers:

Decoding **Section 10 Chemical Quantities Packet Answers**: Revealing the Captivating Potential of Verbal Expression

In an era characterized by interconnectedness and an insatiable thirst for knowledge, the captivating potential of verbal expression has emerged as a formidable force. Its capability to evoke sentiments, stimulate introspection, and incite profound transformations is genuinely awe-inspiring. Within the pages of "**Section 10 Chemical Quantities Packet Answers**," a mesmerizing literary creation penned with a celebrated wordsmith, readers embark on an enlightening odyssey, unraveling the intricate significance of language and its enduring affect our lives. In this appraisal, we shall explore the book is central themes, evaluate its distinctive writing style, and gauge its pervasive influence on the hearts and minds of its readership.

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