

- a) Below, construct an orbital interaction diagram for molecular orbital formation by dragging the images that represent various orbital types (e.g., atomic, bonding, antibonding) into the relevant boxes.
- b) Identify the number of nodes in each atomic and molecular orbital.

Energy ↑

0 nodes
0 nodes
1 node
2 nodes

Select answer

Select answer

Select answer

Continued below...

- c) When two electrons occupy the bonding molecular orbital above, what type of bond results? Explain.

- ☐ σ because only one bond forms.
- ☐ π because the bond is not cylindrically symmetrical about the internuclear axis.
- ☐ σ because the two electrons are in a σ bonding molecular orbital.
- ☐ π because p orbitals can never form σ bonds.

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