AP Chemistry Review Questions

L. Hydrocarbons are acopolar compounds containing carbon and hydrogen atoms. The properties of these hydrocarbons are

ummarized below.	- AND	Electric
Methane	Octans	CIL(CIL), CIL
CH	C _s H _{cs}	Lubricant (grease)
Natural gas	Gasoline	Solid, mp 37 °C
Gas. bp -161 °C	Liquid, bp 126 °C	the same to the Hersi

- st. How do the attractive forces between molecules change in the transition from the gas to the liquid to the solid state?
- A. Based on its proporties, which compound has the strongest attractive forces? The weakest attractive forces?
- c. Write a general statement describing how the size of a molecule influences the strength of London dispersion forces between molecules.
- 2. Dyes are organic compounds that can be used to impart bright, permanent colors to fabrics. The affinity of a dye for a fabric depends on the chemical structures of the dye and fabric molecules and also on the interactions between them. Three common fabrics are wool, cotton and nylon. Wool is a protein, a naturally occurring polymer made up of amino acids with ionized (charged) side chains. Cotton is a naturally occurring polymer made up of glucose units with hydrophilic groups surrounding each glucose unit. Nylon is a synthetic polymer made of hydrocarbon repeating chains joined together by highly polar amide (-CONH-) functional groups.
 - a. The chemical structure of methyl orange is drawn below. Identify the groups in the dye that will bind to ionic and polar sites in a fabric.

& Complete the following "If/then" hypothesis to explain how the structure of a fabric will influence the relative color intensity produced by methyl orange.

"If a fabric contains more ionic and polar groups in its structure, then the intensity of the dye color due to methyl orange should (increase/decrease), because _

c. Using this hypothesis, predict the relative color intensity that would be produced by methyl orange on cotton, nylon and wool. Rank the fabrics from 1 = lightest color to 3 = darkest color.

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