

Chemistry 2nd Semester Final Review **KEY**

This is meant as a guide. It does NOT contain every single thing you need to know for the final exam. Review notes, benchmark review sheets, homeworks, chapter study guides, etc.

Bring your book to school by Monday for turn in (seniors should bring book to bookstore by themselves ASAP)

Gas Laws (Ch. 13):

Subtopics: atmospheric pressure, Dalton's law of partial pressures, combined gas law, Ideal gas law, gas stoichiometry (22.4 L/mol for a gas at STP)

Practice questions:

- If a gas occupies 3.8 L at a pressure of 2.71 atm, what would the pressure be if the volume changes to 1.47 L? (Temp. and amt. of gas constant)
 $V_1 = 3.8 \text{ L}; P_1 = 2.71 \text{ atm}$
 $V_2 = 1.47 \text{ L}; P_2 = ? \text{ Atm}$
Since all units are standard (L, atm), we do not need to convert the units before proceeding.
 $(P_1 V_1)/T_1 = (P_2 V_2)/T_2$
T goes away (not mentioned in problem)
 $P_1 V_1 = P_2 V_2$
Isolate for P_2 : $P_2 = (P_1 V_1)/V_2$
 $P_2 = (2.71 \text{ atm} \times 3.8 \text{ L})/(1.47 \text{ L}) = 7.0 \text{ atm}$ (L cancelled and it's 2 sig figs)
- A gas has a volume of 1.49 L at a temperature of 34.75 degrees Celsius. What would the volume be at 78.41 degrees Celsius? (pressure & amt. of gas constant)
 $V_1 = 1.49 \text{ L}, T_1 = 34.75 \text{ degrees Celsius}$
 $V_2 = ? \text{ L}, T_2 = 78.41 \text{ degrees Celsius}$
We need to convert from Celsius to Kelvins before proceeding.
 $T_1 = 34.75 + 273 = 307.75 \text{ K}$
 $T_2 = 78.41 + 273 = 351.41 \text{ K}$
 $(P_1 V_1)/T_1 = (P_2 V_2)/T_2$
Get rid of P
 $V_1/T_1 = V_2/T_2$
Isolate V_2 : $V_2 = (V_1/T_1)T_2$
 $V_2 = (1.49 \text{ L}/307.75 \text{ K})(351.41 \text{ K}) = 1.70 \text{ L}$
- What volume is occupied by 8.47 g of hydrogen gas at 84.7 degrees Celsius and 1.04 atm?
The combined gas law won't work here. You only have one condition and nothing changes.
 $PV = nRT$ will work. (n = moles and R = 0.08206 L atm/K mol)
 $P = 1.04 \text{ atm}, V = ? \text{ L}, n = \text{OOPS we need to convert from grams to moles!}, T = 84.7 \text{ degrees Celsius} + 273 = 357.7 \text{ K}$
 n : Hydrogen gas is diatomic!
 $8.47 \text{ g H}_2 \times (1 \text{ mol H}_2 / 2.02 \text{ g H}_2) = 4.193069307 \text{ mol H}_2$
Isolate V : $V = (nRT)/P = (4.193069307 \text{ mol H}_2 \times 0.08206 \text{ L atm/K mol} \times 357.7 \text{ K}) / 1.04 \text{ atm} = 118 \text{ L}$
- What volume is occupied by 56.75 g of oxygen gas at STP?

Second Semester Standard Chemistry Review Guide

H Kauffman



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Table of Contents Second Semester Standard Chemistry Review Guide

1. Understanding the eBook Second Semester Standard Chemistry Review Guide
 - The Rise of Digital Reading Second Semester Standard Chemistry Review Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Second Semester Standard Chemistry Review Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Second Semester Standard Chemistry Review Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Second Semester Standard Chemistry Review Guide
 - Personalized Recommendations
 - Second Semester Standard Chemistry Review Guide User Reviews and Ratings
 - Second Semester Standard Chemistry Review Guide and Bestseller Lists

5. Accessing Second Semester Standard Chemistry Review Guide Free and Paid eBooks
 - Second Semester Standard Chemistry Review Guide Public Domain eBooks
 - Second Semester Standard Chemistry Review Guide eBook Subscription Services
 - Second Semester Standard Chemistry Review Guide Budget-Friendly Options
6. Navigating Second Semester Standard Chemistry Review Guide eBook Formats
 - ePub, PDF, MOBI, and More
 - Second Semester Standard Chemistry Review Guide Compatibility with Devices
 - Second Semester Standard Chemistry Review Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Second Semester Standard Chemistry Review Guide
 - Highlighting and Note-Taking Second Semester Standard Chemistry Review Guide
 - Interactive Elements Second Semester Standard Chemistry Review Guide
8. Staying Engaged with Second Semester Standard Chemistry Review Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Second Semester Standard Chemistry Review Guide
9. Balancing eBooks and Physical Books Second Semester Standard Chemistry Review Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Second Semester Standard Chemistry Review Guide
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Second Semester Standard Chemistry Review Guide
 - Setting Reading Goals Second Semester Standard Chemistry Review Guide
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Second Semester Standard Chemistry Review Guide
 - Fact-Checking eBook Content of Second Semester Standard Chemistry Review Guide
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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