

SECOND PUC ANNUAL EXAMINATION APRIL-2006

CHEMISTRY

Time: 3 Hours

Max. Marks: 90

PART - A

NOTE: i) Answer all the 10 questions ii) Questions have to be answered in one word or in one sentence each. Each question carries one mark.

10 X 1 = 10

1. Out of carbon and zinc which is a stronger reducing agent?
2. Why is Fe^{3+} ion more stable than Fe^{2+} ion?
3. If the rate of the reaction is independent of concentration of the reactant, what is the order of the reaction?
4. Which noble gas is not adsorbed by charcoal?
5. What is hydronium ion?
6. What is the only factor on which the colligative property value depends?
7. Between potassium nitrate and aluminium nitrate, which one is required in minimum concentration to coagulate arsenious sulphide sol?
8. Which is the organic product obtained when methyl magnesium iodide reacts with formaldehyde followed by acid hydrolysis?
9. Name the least stable cycloalkane.
10. How many amino acid units are present in insulin?

PART - B

NOTE: i) Answer any ten questions. ii) Each question carries two marks.

10 X 2 = 20

11. Justify that concentrated sulphuric acid is a good dehydrating agent with example.
12. Calculate the bond order in oxygen molecule.
13. The time for half-life of a first order reaction found to be 30 mins. What is the rate constant of the reaction?
14. Define Faraday's first law of electrolysis.
15. What is a spontaneous reaction? Explain with an example.
16. What is Tyndall effect? Explain.
17. Calculate the number of particles present in the unit cell of BCC.
18. Distinguish between inductive effect and electromeric effect.
19. Explain the effect of the following substituents on the acidity of phenol:
i) $-\text{CH}_3$ group ii) $-\text{NO}_2$ group.
20. How is ethanamine prepared from ethane nitrile?
21. What are reducing and non-reducing sugars? Give one example for each.
22. What are biological messengers? Give an example.

PART - C

I. Answer any two of the following:

2 X 5 = 10

23. a) How are noble gases separated by Dewar's charcoal adsorption method?
b) With the help of the Ellingham's diagram, explain why silver oxide can be thermodynamically decomposed at relatively lower temperature.
24. a) Give the postulates of Werner's theory of coordination compounds.
b) Copper (I) compounds are white and diamagnetic whereas Copper (II) compounds are coloured and paramagnetic. Give reason.
25. a) Describe the manufacture of sodium hydroxide using Nelson's cell.
b) Write the IUPAC name of $[\text{Co}(\text{NH}_3)_4\text{Cl}_2]\text{Cl}$

II. Answer any three of the following:

3 X 5 = 15

26. a) Discuss the molecular orbital structure of Benzene.
b) Explain geometrical isomerism with an example.
27. a) What is $\text{S}_\text{N}1$ mechanism? Explain it using hydrolysis of t-butyl bromide.
b) Give any two uses of Aniline.
28. a) Write the Haworth's ring structure for $\beta\text{-D}(-)\text{fructopyranose}$.
b) Define the term 'oxidation value of an oil'. Mention its significance.

Second Puc Question Paper Chemistry march 2006

Camilla Rothe



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