

Periodic Table Questions and Answers

How many elements are there?

118

What is the periodic table?

A chart that organizes elements by atomic number and shared properties

Who invented the periodic table?

Dmitri Mendeleev

A chart that organizes elements by atomic number and shared properties

The periodic table is displayed with elements color-coded by groups. The noble gases (Group 18) are highlighted in yellow. The lanthanide and actinide series are shown at the bottom. The table includes element symbols, atomic numbers, and names.

1 H Hydrogen	2 He Helium																	11 B Boron	12 C Carbon	13 N Nitrogen	14 O Oxygen	15 F Fluorine	16 Ne Neon
3 Li Lithium	4 Be Beryllium																	17 Al Aluminum	18 Si Silicon	19 P Phosphorus	20 S Sulfur	21 Cl Chlorine	22 Ar Argon
5 Na Sodium	6 Mg Magnesium	7 Sc Scandium	8 Ti Titanium	9 V Vanadium	10 Cr Chromium	11 Mn Manganese	12 Fe Iron	13 Co Cobalt	14 Ni Nickel	15 Cu Copper	16 Zn Zinc	17 Ga Gallium	18 Ge Germanium	19 As Arsenic	20 Se Selenium	21 Br Bromine	22 Kr Krypton						
23 K Potassium	24 Ca Calcium	25 Sc Scandium	26 Ti Titanium	27 V Vanadium	28 Cr Chromium	29 Mn Manganese	30 Fe Iron	31 Co Cobalt	32 Ni Nickel	33 Cu Copper	34 Zn Zinc	35 Ga Gallium	36 Ge Germanium	37 As Arsenic	38 Se Selenium	39 Br Bromine	40 Kr Krypton						
39 Rb Rubidium	40 Sr Strontium	41 Y Yttrium	42 Zr Zirconium	43 Nb Niobium	44 Mo Molybdenum	45 Tc Technetium	46 Ru Ruthenium	47 Rh Rhodium	48 Pd Palladium	49 Ag Silver	50 Cd Cadmium	51 In Indium	52 Sn Tin	53 Sb Antimony	54 Te Tellurium	55 I Iodine	56 Xe Xenon						
55 Cs Cesium	56 Ba Barium	57 La Lanthanum	58 Ce Cerium	59 Pr Praseodymium	60 Nd Neodymium	61 Pm Promethium	62 Sm Samarium	63 Eu Europium	64 Gd Gadolinium	65 Tb Terbium	66 Dy Dysprosium	67 Ho Holmium	68 Er Erbium	69 Tm Thulium	70 Y Ytterbium	71 Lu Lutetium	72 Hf Hafnium						
73 Fr Francium	74 Ra Radium	75 Ac Actinium	76 Th Thorium	77 Pa Protactinium	78 U Uranium	79 Np Neptunium	80 Pu Plutonium	81 Am Americium	82 Cm Curium	83 Bk Berkelium	84 Cf Californium	85 Es Einsteinium	86 Fm Fermium	87 Md Mendelevium	88 No Nobelium	89 Lr Lawrencium	90 Ug Ununseptium						

periodic
nds?
r, ionization

Why a
They have

What are the periodic table trends?

electronegativity, ionization energy, atomic radius, electron affinity

Why are noble gas so inert?

They have stable electron configurations.

Section 2 Exploring The Periodic Table Answers

Kaplan Test Prep



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This full color manual is designed to satisfy the content needs of either a one or two semester introduction to physical science course populated by nonmajors It provides students with the opportunity to explore and make sense of the world around them to develop their skills and knowledge and to learn to think like scientists The material is written in an accessible way providing clearly written procedures a wide variety of exercises from which instructors can choose and real world examples that keep the content engaging Exploring Physical Science in the Laboratory guides students through the mysteries of the observable world and helps them develop a clear understanding of challenging concepts **Chemistry II For Dummies**

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Worksheets, Student Edition, for Use with Glencoe Physical Science Aron Thompson, 1999 **Decoding the**

Periodic Table Jurjen Van Van der Wal,2010-06-16 Decoding the Periodic Table In the year 1809 Russian Chemist Dimitri Ivanovich Mendeleev published the first Chemical Periodic Table of the Elements Since then scientists developed a theory that every element consists of a tightly packed mass of protons and neutrons which is surrounded by concentric spheres which are occupied by orbiting electrons This concept is known as the Standard Model but it is flawed because it cannot accommodate gravity For many decades researchers including this author in the late 1960 s tried to formulate theories that in a comprehensive way would unite all physical designs and phenomena under a common denominator but to no avail The String Theory is one of them but it has found no success In 1989 the author of the book made a fresh start again with the notion that particles such as the proton neutron and electron would not be round but that they would have straight lined square or cubic configurations because the Periodic Table s had quadratic numbers in their quantities of elements such as Period 1 with 2 elements 2 2x12 Periods 2 3 8 2 x 22 Periods 4 5 18 2 x 32 and Period 6 7 32 2 x 42 This might indicate square structural relationships lets give it a try It worked it took 20 years to do it it was a fascinating journey falling down and getting up again and again working without a computer with all figures hand drawn and discovering that the forces that make the nuclear mass of the neutron do that by multiplying two force filled cubic volumes as in $3F \cdot 6 \cdot 3F \cdot 3F \cdot 3 \cdot 3F \cdot 3F \cdot 3F$ 1838 2656

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