

Science

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(Chapter – 1) (Chemical Reactions and Equations) (Intext)
(Class 10)

Page 10

Question 1:

A solution of a substance 'X' is used for white washing.

(i) Name the substance 'X' and write its formula.

(ii) Write the reaction of the substance 'X' named in (i) above with water.

Answer 1:

(i) The substance 'X' is calcium oxide. Its chemical formula is CaO.

(ii) Calcium oxide reacts vigorously with water to form calcium hydroxide (slaked lime)



Question 2:

Why is the amount of gas collected in one of the test tubes in Activity 1.7 double of the amount collected in the other? Name this gas.

Answer 2:

During the Electrolysis of water, hydrogen and oxygen is get separated by the electricity. Water (H₂O) contains two parts hydrogen and one part oxygen. Since hydrogen goes to one test tube and oxygen goes to another, the amount of gas collected in one of the test tubes is double of the amount collected in the other.

Page 13

Question 1:

Why does the colour of copper sulphate solution change when an iron nail is dipped in it?

Answer 1:

When an iron nail is dipped in a copper sulphate solution, iron (which is more reactive than copper) displaces copper from copper sulphate solution forming iron sulphate, which is green in colour.



Therefore, the blue colour of copper sulphate solution fades and green colour appears.

Question 2:

Give an example of a double displacement reaction other than the one given in Activity 1.10.

Answer 2:

Sodium carbonate reacts with calcium chloride to form calcium carbonate and sodium chloride.



In this reaction, sodium carbonate and calcium chloride exchange ions to form two new compounds. Hence, it is a double displacement reaction.

Question 3:

Identify the substances that are oxidised and the substances that are reduced in the following reactions.

(i) $4\text{Na(s)} + \text{O}_2\text{(g)} \rightarrow 2\text{Na}_2\text{O(s)}$

(ii) $\text{CuO(s)} + \text{H}_2\text{(g)} \rightarrow \text{Cu(s)} + \text{H}_2\text{O(l)}$

Answer 3:

(i) Sodium (Na) is oxidised as it gains oxygen and oxygen gets reduced.

(ii) Copper oxide (CuO) is reduced to copper (Cu) while hydrogen (H₂) gets oxidised to water (H₂O).

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Lord Eccles, Wooster



Section 10 1 Reactions And Equations Answers Pg 55:

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Table of Contents Section 10 1 Reactions And Equations Answers Pg 55

1. Understanding the eBook Section 10 1 Reactions And Equations Answers Pg 55
 - The Rise of Digital Reading Section 10 1 Reactions And Equations Answers Pg 55
 - Advantages of eBooks Over Traditional Books
2. Identifying Section 10 1 Reactions And Equations Answers Pg 55
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Section 10 1 Reactions And Equations Answers Pg 55
 - User-Friendly Interface
4. Exploring eBook Recommendations from Section 10 1 Reactions And Equations Answers Pg 55
 - Personalized Recommendations
 - Section 10 1 Reactions And Equations Answers Pg 55 User Reviews and Ratings
 - Section 10 1 Reactions And Equations Answers Pg 55 and Bestseller Lists
5. Accessing Section 10 1 Reactions And Equations Answers Pg 55 Free and Paid eBooks
 - Section 10 1 Reactions And Equations Answers Pg 55 Public Domain eBooks
 - Section 10 1 Reactions And Equations Answers Pg 55 eBook Subscription Services
 - Section 10 1 Reactions And Equations Answers Pg 55 Budget-Friendly Options
6. Navigating Section 10 1 Reactions And Equations Answers Pg 55 eBook Formats

- ePub, PDF, MOBI, and More
 - Section 10 1 Reactions And Equations Answers Pg 55 Compatibility with Devices
 - Section 10 1 Reactions And Equations Answers Pg 55 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Section 10 1 Reactions And Equations Answers Pg 55
 - Highlighting and Note-Taking Section 10 1 Reactions And Equations Answers Pg 55
 - Interactive Elements Section 10 1 Reactions And Equations Answers Pg 55
 8. Staying Engaged with Section 10 1 Reactions And Equations Answers Pg 55
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Section 10 1 Reactions And Equations Answers Pg 55
 9. Balancing eBooks and Physical Books Section 10 1 Reactions And Equations Answers Pg 55
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Section 10 1 Reactions And Equations Answers Pg 55
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Section 10 1 Reactions And Equations Answers Pg 55
 - Setting Reading Goals Section 10 1 Reactions And Equations Answers Pg 55
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Section 10 1 Reactions And Equations Answers Pg 55
 - Fact-Checking eBook Content of Section 10 1 Reactions And Equations Answers Pg 55
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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