| Name:            |  |
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# **Nuclear Decay**

Types of Radiation

Complete the table below to compare the properties of alpha, beta, positron, and gamma radiation.

| Name         | Alpha | Beta | Positron | Gamma |
|--------------|-------|------|----------|-------|
| Greek Letter |       |      |          |       |
| Symbol       |       | 1    |          |       |
| Charge       |       |      |          |       |
| Stopped by   |       |      |          |       |

Nuclear Decay Reactions

In problem 2-5, fill in the missing particle for each nuclear decay reaction. Then, match each equation to the appropriate type of nuclear decay.

A. alpha emission

B. beta emission

C. electron capture

D. positron emission

In problem 6 and 7, write a balanced equation for each for each nuclear decay reaction.

- 6. Decay of Radium-223 by alpha emission 223<sub>88</sub> Ra - 4<sub>2</sub> He +

7. Decay of Iodine-126 by positron emission 126 st 1 - 0 +1 e +

8. How are the atomic number and mass number of a nuclide affected by positron emission?

9. How are the atomic number and mass number of a nuclide affected by the emission of gamma radiation?

10. Which has a slightly greater mass in a nuclear reaction, the reactants or the product? Why?

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