



Lab Notebook

Classifying Chemical Reactions

Write the balanced chemical equation for each reaction described in the boxes. Identify the reaction type for each reaction. Write the name of the reaction in the box.

Column I

1. _____ A. A precipitate forms in a gas liquid collection over water (combustion, decomposition, precipitation)
2. _____ B. Two reactants exchange products forming another substance
3. _____ C. A hydrocarbon explodes, creating an inorganic
4. _____ D. Chloroalkanes break down into hydrogen and alkenes
5. _____ E. A type of redox reaction that produces heat

Write the balanced chemical equation for each reaction described in the space provided.

combustion _____ A. $\text{CH}_4 + 2\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$

precipitation _____ B. $\text{NaCl} + \text{AgNO}_3 \rightarrow \text{NaNO}_3 + \text{AgCl}$

decomposition

synthesis reaction

_____ C. $\text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$

single displacement reaction _____ D. $\text{Zn} + \text{CuSO}_4 \rightarrow \text{ZnSO}_4 + \text{Cu}$

synthesis reaction

_____ E. $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$

_____ F. $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$

combustion

single displacement reaction

_____ G. $\text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$

synthesis reaction

single displacement reaction

_____ H. $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$

combustion

single displacement reaction

_____ I. $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$

single displacement reaction _____ J. $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$

single displacement reaction

Section 2 Reinforcement Classifying Chemical Reactions

RJ Alexander



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