

Answers: Rates of Reactions Practice QuestionsRate Laws:

1.



Initial Concentration (mol/L)		Initial Rate (mol/L.s)
$\text{NO}_{(\text{g})}$	$\text{H}_{2(\text{g})}$	
0.075	0.400	0.005
0.150	0.200	0.010
0.150	0.400	0.020

- a) Find the rate law. **Rate = $k[\text{NO}]^2[\text{H}_2]$**
 b) Determine the value of the rate constant.
 $0.005 = k(0.075)^2(0.400)$
 $k = 2.22 \text{ L}^2/\text{mol}^2\cdot\text{s}$

2.



Initial Concentration (mol/L)		Initial Rate (mol/L.s)
$\text{NO}_2^-_{(\text{aq})}$	$\text{NH}_4^+_{(\text{aq})}$	
0.010	0.020	5.4×10^{-7}
0.020	0.020	10.8×10^{-7}
0.040	0.020	21.6×10^{-7}
0.020	0.0202	10.8×10^{-7}
0.020	0.0404	21.6×10^{-7}
0.020	0.0606	32.4×10^{-7}

- a) Write the rate law. **Rate = $k[\text{NO}_2][\text{NH}_4^+]$**
 b) Calculate the rate constant.
 $5.4 \times 10^{-7} = k(0.01)(0.02)$
 $k = 0.0027 \text{ L/mol.s}$
 c) Calculate the initial rate if $[\text{NO}_2^-_{(\text{aq})}] = 0.10 \text{ M}$ and $[\text{NH}_4^+_{(\text{aq})}] = 0.10 \text{ M}$.
Rate = $0.0027(0.10)(0.10)$
Rate = $2.7 \times 10^{-5} \text{ mol/L.s}$

3.



Initial Concentration (mol/L)		Rate of Disappearance of H_2S (mol/L.s)
$\text{OH}_{(\text{g})}$	$\text{H}_2\text{S}_{(\text{g})}$	
1.3×10^{-3}	2.1×10^{-3}	1.4×10^{-5}
3.9×10^{-3}	2.1×10^{-3}	4.2×10^{-5}
3.9×10^{-3}	4.2×10^{-3}	8.4×10^{-5}

Section 18 Rates Of Reaction 467 Answers

Lars I. Eriksson



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