

## Practice Problems

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### SECTION 12.1 THE MOLE: A MEASUREMENT OF MATTER

1. Calculate the molar mass of each of the following compounds:
- phosphorus pentachloride ( $\text{PCl}_5$ )
  - carbon tetrachloride ( $\text{CCl}_4$ )
2. Calculate the mass percent of each element in the compound:
- $\text{KNO}_3$
  - $\text{Fe}_2\text{SO}_4$
3. Determine the number of moles of each element in 1.000 g of:
- hydrogen peroxide ( $\text{H}_2\text{O}_2$ )
  - acetic acid ( $\text{CH}_3\text{COOH}$ )
4. Determine the mass of 1.00 mole of each of the following:
- hydrogen peroxide ( $\text{H}_2\text{O}_2$ )
  - acetic acid ( $\text{CH}_3\text{COOH}$ )

### SECTION 12.2 MOLE–MASS AND MOLE–VOLUME RELATIONSHIPS

1. Determine the molar mass of each of the following compounds:
- $\text{H}_2\text{O}$
  - $\text{H}_2\text{SO}_4$
  - $\text{CO}_2$
  - $\text{C}_2\text{H}_6$
2. Calculate the mass in grams of each of the following:
- 0.500 mole of water ( $\text{H}_2\text{O}$ )
  - 0.250 mole of carbon dioxide ( $\text{CO}_2$ )
  - 0.100 mole of hydrogen gas ( $\text{H}_2$ )
  - 0.050 mole of oxygen gas ( $\text{O}_2$ )
3. Determine the mass in grams of each of the following:
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  - 0.100 mole of hydrogen gas ( $\text{H}_2$ )
  - 0.050 mole of oxygen gas ( $\text{O}_2$ )

# Section 10 Chemical Quantities Practice Problem Answers

**Victor CM Freestone**



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system designed to help you pass the FE exam the first time Exam Topics Covered Chemical Reaction Chemistry Computational Tools Engineering Engineering Sciences Ethics and Professional Practice Fluid Mechanics Dynamics Heat Transfer Mass Transfer and Separation Material Energy Balances Materials Science Mathematics Probability and Statistics Process Control Process Design and Economics Safety Health and Environment Thermodynamics Key Features Over 600 three minute multiple choice exam like practice problems to illustrate the type of problems you ll encounter during the exam Clear complete and easy to follow solutions to deepen your understanding of all knowledge areas covered in the exam Step by step calculations using equations and nomenclature from the NCEES FE Reference Handbook to familiarize you with the reference you ll have on exam day Binding Paperback Publisher PPI A Kaplan Company [Chemical Surety](#) United States. Department of the Army,1995 *Elementary Principles of Chemical Processes* Richard M. Felder,Ronald W. Rousseau,Lisa G. Bullard,2020-08-11 This best selling text prepares students to formulate and solve material and energy balances in chemical process systems and lays the foundation for subsequent courses in chemical engineering The text provides a realistic informative and positive introduction to the practice of chemical engineering **Crystal Structures** M Ladd,1999-06 This book presents and discusses those common crystal structures that would be encountered by students taking chemistry or any subject within which chemistry forms a significant component With many worked examples and a wide selection of problems with solutions Includes instructions for making simple stereoviewers and computer programs in a thorough treatment of binary alloys and three dimensional packing in molecular solids [Chemical Thermodynamics](#) Victor CM Freestone,2025-06-13 Thermodynamics can never be made easy but with the right approach and a consistent use of scientific terms it can be made less opaque and it can give a person who is prepared to try an insight into how science explains why things happen the way they do The approach adopted in this book will give readers a better understanding of how science works together with its limitations Unfortunately thermodynamics or at least some parts of it is a subject which apart from quantum mechanics probably causes most confusion and bewilderment amongst scientists The majority of students do not understand or get thermodynamics and it is considered a hard or difficult subject There are multiple reasons for this There is of course mathematics and many thermodynamic texts appear to be lists upon lists of differential equations Another reason is that thermodynamics is as often as not poorly taught by teachers lecturers who themselves do not understand or appreciate or have any interest in the subject often all three This results not only in a lack of scientific rigorousness in the teaching of the subject with the resulting confusion and sometimes teachers lecturers and authors just get it plain wrong this occurs surprisingly often However it need not be like this and although mathematics including calculus is required it can be kept to a relatively elementary level in order to obtain an understanding of this most important of subjects No one can pretend that the subject is easy but it can be made more accessible by a rigorous definition of terms and concepts and ensuring that a consistency of use of these definitions is maintained Highlighting the benefits of

thermodynamics in practical science the text gives an intuitive grasp of the major concepts of thermodynamics such as energy and entropy Provides a new pedagogic approach to understanding and teaching chemical thermodynamics Starting with a set of basic simple assumptions about what constitutes topics such as an ideal gas theories are developed in a clear concise and accessible manner that will either answer or at the very least give an insight into a surprising range of scientific phenomena including energy heat temperature properties of gases time and quantum theory Assumes that the reader has essentially no knowledge of the subject Mathematics including calculus is kept to a relatively elementary level in order to obtain an understanding of this most important of subjects Provides the reader with a better understanding of how science works together with its limitations

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Harry L. Pardue, 2018-11-15 Concepts procedures and programs described in this book make it possible for readers to solve both simple and complex equilibria problems quickly and easily and to visualize results in both numerical and graphical forms They allow the user to calculate concentrations of reactants and products for both simple and complicated situations The user can spend less time doing calculations and more time thinking about what the results mean in terms of a larger problem in which she or he may be interested

**Food Analysis** Suzanne Nielsen, 2003-04-30 This book provides information on the techniques needed to analyze foods in laboratory experiments All topics covered include information on the basic principles procedures advantages limitations and applications This book is ideal for undergraduate courses in food analysis and is also an invaluable reference to professionals in the food industry General information is provided on regulations standards labeling sampling and data handling as background for chapters on specific methods to determine the chemical composition and characteristics of foods Large expanded sections on spectroscopy and chromatography are also

included Other methods and instrumentation such as thermal analysis selective electrodes enzymes and immunoassays are covered from the perspective of their use in the chemical analysis of foods A helpful Instructor s Manual is available to adopting professors

**Chemistry** John A. Olmsted, Gregory M. Williams, Robert Charles Burk, 2016-01-14 Olmsted Burk is an introductory general chemistry text designed specifically with Canadian professors and students in mind A reorganized Table of Contents and inclusion of SI units IUPAC standards and Canadian content designed to engage and motivate readers distinguish this text from many of the current text offerings It more accurately reflects the curriculum of most Canadian institutions Instructors will find the text sufficiently rigorous while it engages and retains student interest through its accessible language and clear problem solving program without an excess of material that makes most text appear daunting and redundant

**Insights into Chemical Engineering** P. V. Danckwerts, 2013-10-22 A selection of papers many of which proved novel and thought provoking and have had a considerable influence on the development of chemical engineering chosen by Professor Danckwerts from research work conducted at Cambridge and Imperial College mainly during the years 1950 1954 and 1957 1973 They are divided into 6 sections with linking critical commentaries

**Introduction to Chemical Science** Mary Maier, 1978

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*Pharmaceutical and Chemical Problems and Exercises* Oscar Oldberg, 1907

**Investigating Chemistry** Matthew Johll, 2008-12-22 In its new second edition Investigating Chemistry A Forensic Science Perspective remains the only book that uses the inherently fascinating topics of crime and criminal investigations as a context for teaching the fundamental chemical concepts most often covered in an introductory nonmajors course Covering all the standard topics Matthew Johll capitalizes on the surge of interest in the scientific investigation of crime as sparked by CSI and other television shows bringing together the theme of forensic science and the fundamentals of chemistry in ways that are effective and accessible for students This edition features refined explanations of the chemical concepts which are the core of the book as well as a more thoroughly integrated forensic theme updated features and an expanded media supplements package

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