Name Date Class

Chemical Reactions . Review and Reinforce

# **Describing Chemical Reactions**

#### **Understanding Main Ideas**

Balance the equations on the lines below. State whether the reaction is a synthesis, decomposition, or replacement reaction.

| Given Equation  | Balanced Equation | Type of Reaction |
|---|-------------------|------------------|
| $f_* \text{ FeS} * \text{ HCL} \rightarrow \text{FeCl}_2 * \text{ H}_2\text{S}$ | w.                | ъ.               |
| 2. Na + F <sub>2</sub> → NaF  | +.                | b.               |
| 3. HgO → H <sub>S</sub> + O <sub>2</sub>  | e.                | ъ.               |

Answer questions 4 and 5 on a separate sheet of paper.

- Describe in words the chemical composition of the molecules involved and the reaction represented by the equation: 2H<sub>2</sub> + O<sub>2</sub> → 2 H<sub>2</sub>O
- Use the principle of conservation of mass to explain why the equation in question 6 is balanced.

#### **Building Vocabulary**

Match each term with its difficultion by writing the letter of the correct difficultion in the right column on the line beside the term in the lift column.

| 6. chemical equation      | <ul> <li>a. substance present after a reaction.</li> </ul>   |
|---------------------------|--|
| 7. decomposition reaction | <ul> <li>b. reaction in which substances combine to<br/>form a more complex compound</li> </ul>                    |
| B. coefficient            | <ul> <li>uses symbols and formulas to show<br/>cheretoal reactions</li> </ul>                                      |
| 9. product                | <ul> <li>d. reaction in which one element replaces<br/>another in a compound.</li> </ul>                           |
| 16 montant                | <ul> <li>substance present before a reaction</li> </ul>  |
| 11. conservation of mass  | <ol> <li>rumber telling hose many molecules<br/>of a substance are insolved in a chemical<br/>reaction.</li> </ol> |
| 12. synthesis maction     | <ul> <li>g. reaction in which compounds are broken<br/>down into simpler products</li> </ul>                       |
| 13- replacement reaction  | <ul> <li>b. principle that states that matter is not<br/>created or destroyed during a chemical</li> </ul>         |

resoltion.

# **Section 1 2 Describing Chemical Reactions Review And Reinforce**

Joacim Rocklöv

#### **Section 1 2 Describing Chemical Reactions Review And Reinforce:**

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