

Physical Science – 2nd Semester – Final Exam Study Guide – (2012–2013)

1. Physical Science Basics

- What tool(s) would you want to use to find the:
 - Mass of an object? triple beam balance
 - Volume of an object? ruler or water displacement method
 - Length of an object? Ruler
 - Density of an object? graduated cylinder
- What is the formula for calculating density? mass/volume
 - What is the density of a cube that is 33 g and has a volume of 3 cm³? 11g/cm³
 - Compare the density of solid H₂O to liquid H₂O, and explain what observable effect this has.
solid H₂O less dense than liquid H₂O

2. Chemistry Basics

- Element (def.) substance made up of one kind of atom
- Atom (def.) basic building block of matter
- Compound (def.) a molecule that contains atoms or molecules of two or more different elements
- Molecule (def.) a combination of two or more of the same or different atoms that are bonded together

3. Model of the Atomic Evolution of the Atomic Theory

Atoms are so small that many scientists throughout history have created models to describe them.

Why have atomic theories changed over time? Better technology came around.

Scientist	Year	Major Discovery/ Contribution	Name of atomic model
John Dalton	1808	Each element is made up of small, indivisible particles called atoms. Different elements have atoms of different masses.	Billiard Ball Model
J.J. Thomson	1897	Negatively charged electron	Plum Pudding Model
Niels Bohr	1913	Electrons move <u>fixed orbits</u> around the atom's nucleus in specific layers, or <u>electron shells</u> . Atoms absorb or give off energy when the electrons move from one shell (<u>energy level</u>) to another.	Bohr Model
Ernest Rutherford	1909	1909 /Based on idea that positive reflects positive/ atom mostly empty space; electrons orbit randomly around positively charged nucleus	Planetary Model

4. Periodic Table

- The current periodic table is arranged by increasing atomic number.
- Reading the Periodic Table
 - An element's properties can be predicted from its location in the periodic table.
 - The atomic number represents the number of protons.
 - The atomic mass represents the number of protons and neutrons.

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A red circular graphic with a gradient, appearing as a partial circle or a thick arc, located to the right of the text box.

Second Semester Final Study Guide:

, *A Manual of the Wausau Public Schools* Wausau (Wis.). Board of Education, 1914

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The Mobile Learning Voyage - From Small Ripples to Massive Open Waters Tom H. Brown, Herman J. van der Merwe, 2015-10-14 This book constitutes the proceedings of the 14th International Conference on Mobile and Contextual Learning mLearn 2015 held in a cruise ship leaving from and arriving to Venice Italy in October 2015 The 22 revised full papers and 6 short papers presented were carefully reviewed and selected from 81 submissions The papers deal with the topics related to the theme of the conference The mobile learning voyage from small ripples to massive open waters The conference theme paid tribute to the developments that brought mobile learning from its infancy steps in the early 2000s to maturity in 2015 while simultaneously paving the way for the broad and open waters ahead with new developments and progress in mobile learning and emerging ambient technologies

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