CHAPTER 9 REVIEW

Stoichiometry

SECTION 1

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- The coefficients in a chemical equation represent the
 - (a) masses in grams of all reactants and products.
 - (b) relative number of moles of reactants and products.
 - (c) number of atoms of each element in each compound in a reaction.
 - (d) number of valence electrons involved in a reaction.
- Which of the following would not be studied within the topic of stoichiometry?
 - (a) the mole ratio of Al to Cl in the compound aluminum chloride
 - (b) the mass of carbon produced when a known mass of sucrose decomposes
 - (c) the number of moles of hydrogen that will react with a known quantity of oxygen
 - (d) the amount of energy required to break the ionic bonds in CaF2
- A balanced chemical equation allows you to determine the
 - (a) mole ratio of any two substances in the reaction.
 - (b) energy released in the reaction.
 - (c) electron configuration of all elements in the reaction.
 - (d) reaction mechanism involved in the reaction.
- The relative number of moles of hydrogen to moles of oxygen that react to form water represents a(n)
 - (a) reaction sequence.
 - (b) bond energy.
 - (c) mole ratio.
 - (d) element proportion.
- 5. Given the reaction represented by the following unbalanced equation:

 $N_2O(g) + O_2(g) \rightarrow NO_2(g)$

a. Balance the equation.

b. What is the mole ratio of NO2 to O2?

- c. If 20.0 mol of NO₂ form, how many moles of O₂ must have been consumed?
 - d. Twice as many moles of NO₂ form as moles of N₂O are consumed. True or False?
- e. Twice as many grams of NO₂ form as grams of N₂O are consumed. True or False?

Review Stoichiometry Section 1 Short Answer

Kenneth S. Overway

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