

Model 2 – Ternary Acids (Oxyacids)

Acid	Name of Acid in Aqueous Solution	Cation (+)	Polyatomic Anion (-)	Polyatomic Anion Name
HClO_3	Chloric acid	H_3O^+	ClO_3^{-1}	Chlorate
H_2SO_3	Sulfurous acid	$2\text{H}_3\text{O}^+$	SO_3^{-2}	Sulfite
H_2SO_4	Sulfuric acid	$2\text{H}_3\text{O}^+$	SO_4^{-2}	Sulfate
H_3PO_3	Phosphorous acid	$3\text{H}_3\text{O}^+$	PO_3^{-3}	Phosphite
H_3PO_4	Phosphoric acid	$3\text{H}_3\text{O}^+$	PO_4^{-3}	Phosphate
HNO_3	Nitric acid	H_3O^+	NO_3^{-1}	Nitrate
HNO_2	Nitrous acid	H_3O^+	NO_2^{-1}	Nitrite
H_2CO_3	Carbonic acid	$2\text{H}_3\text{O}^+$	CO_3^{-2}	Carbonate

6. Look at the formulas of the ternary acids in Model 2.

a. How are ternary acids different from binary acids in their structure?

Ternary acids contain three elements and they also contain the element oxygen.

b. What number do you think the prefix "ter-" refers to?

"ter" refers to three

7. When ternary acids are mixed with water, ions will form. Fill in the table above with the formulas and names of the anions.

8. Examine the pairs of ternary acids in Model 2 that contain sulfur, phosphorus, and nitrogen. Each pair has one acid that ends in "-ic" and another that ends in "-ous." These endings are related to the name of the polyatomic anion found in the acid ("-ate" or "-ite"). Complete the statements below with the correct acid name ending.

Polyatomic anion ending is "-ate" → acid name ending is "ic"

Polyatomic anion ending is "-ite" → acid name ending is "ous"

9. If the prefix "hydro-" were used to name a ternary acid, what problem would this create when naming HClO_3 ?

ClO_3^{-1} has the name of Chlorate. "ate" must be replaced with "ic". If "hydro" is used as a prefix, it would be hydrochloric acid. This has the formula HCl , not HClO_3 .

10. Write a rule for naming ternary acids.

- The prefix "hydro" is not used in the name when an acid contains more than two elements and one of the elements is oxygen.

- "ate" is replaced with "ic" and "ite" is replaced with "ous".

11. Predict the formula for chlorous acid.



12. Circle the acid(s) below that would be named beginning with the prefix "hydro-."



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Table of Contents Polyatomic Ions Worksheet Pogil

1. Understanding the eBook Polyatomic Ions Worksheet Pogil
 - The Rise of Digital Reading Polyatomic Ions Worksheet Pogil
 - Advantages of eBooks Over Traditional Books
2. Identifying Polyatomic Ions Worksheet Pogil
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Polyatomic Ions Worksheet Pogil
 - User-Friendly Interface
4. Exploring eBook Recommendations from Polyatomic Ions Worksheet Pogil

- Personalized Recommendations
- Polyatomic Ions Worksheet Pogil User Reviews and Ratings
- Polyatomic Ions Worksheet Pogil and Bestseller Lists
- 5. Accessing Polyatomic Ions Worksheet Pogil Free and Paid eBooks
 - Polyatomic Ions Worksheet Pogil Public Domain eBooks
 - Polyatomic Ions Worksheet Pogil eBook Subscription Services
 - Polyatomic Ions Worksheet Pogil Budget-Friendly Options
- 6. Navigating Polyatomic Ions Worksheet Pogil eBook Formats
 - ePub, PDF, MOBI, and More
 - Polyatomic Ions Worksheet Pogil Compatibility with Devices
 - Polyatomic Ions Worksheet Pogil Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Polyatomic Ions Worksheet Pogil
 - Highlighting and Note-Taking Polyatomic Ions Worksheet Pogil
 - Interactive Elements Polyatomic Ions Worksheet Pogil
- 8. Staying Engaged with Polyatomic Ions Worksheet Pogil
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Polyatomic Ions Worksheet Pogil
- 9. Balancing eBooks and Physical Books Polyatomic Ions Worksheet Pogil
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Polyatomic Ions Worksheet Pogil
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Polyatomic Ions Worksheet Pogil
 - Setting Reading Goals Polyatomic Ions Worksheet Pogil
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Polyatomic Ions Worksheet Pogil

- Fact-Checking eBook Content of Polyatomic Ions Worksheet Pogil
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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