

Flame Test Lab

Activity Key

Background

The heat from a laboratory burner will cause the ions of some elements to give off light. Electrons will absorb the heat energy from the flame and will “jump” to a higher energy level. When the electrons return to their original energy levels, this absorbed energy is released as light. Different elements absorb and release different amounts of energy and thus produce different kinds of light. This light may be visible or invisible, infrared or ultraviolet radiation. Analysis of this light can be used to identify elements and molecules. In today’s lab, we will observe the visible range of light emission.

Objectives

1. Perform a flame test to identify the characteristic color of metal ions.
2. Identify the metal ion in an unknown solution.
3. Calculate the energy of emitted photons.

Materials: (per lab group)

10 ml dilute solutions of the following:

Barium Chloride
Calcium Chloride
Lithium Chloride
Potassium Chloride
Sodium Chloride
Strontium Chloride
Unknown Solution

7 Wood Splints

1 Bunsen Burner

Safety Goggles & Aprons

Note:

If chloride compounds are not available, metal nitrate compounds may be substituted.

Use dilute or approximately 0.1 M solutions

Unknowns:

Number the beakers 1-12.

Add solutions as follows:

1. Strontium Chloride

2. Calcium Chloride

3. Potassium Chloride

4. Barium Chloride

5. Lithium Chloride

6. Sodium Chloride

7. Calcium Chloride

8. Strontium Chloride

9. Barium Chloride

10. Potassium Chloride

11. Sodium Chloride

12. Lithium Chloride

Procedure

1. Put on your safety goggles and an apron.
2. Obtain a wood splint and the barium chloride solution.
3. Dip the wood splint into the solution and allow it to soak for 1-2 minutes.
4. Light Bunsen Burner. USE CAUTION!
5. Remove the splint from the solution and hold the end of the splint into the flame. Do not burn the splint.
6. Observe and record the color of the flame.
7. Repeat steps 2-6 using the remaining chloride solutions.
8. Obtain an unknown solution from your teacher. Record the number of your unknown.
9. Soak a wood splint in the unknown solution.
10. Hold the splint in the flame and record the color of the flame that is produced.
11. Using your data, identify the metal ion in your unknown solution.

Prentice Hall Flame Test Lab Answers

Lauren Gardner



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Table of Contents Prentice Hall Flame Test Lab Answers

1. Understanding the eBook Prentice Hall Flame Test Lab Answers
 - The Rise of Digital Reading Prentice Hall Flame Test Lab Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Prentice Hall Flame Test Lab Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Prentice Hall Flame Test Lab Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Prentice Hall Flame Test Lab Answers
 - Personalized Recommendations
 - Prentice Hall Flame Test Lab Answers User Reviews and Ratings
 - Prentice Hall Flame Test Lab Answers and Bestseller Lists
5. Accessing Prentice Hall Flame Test Lab Answers Free and Paid eBooks
 - Prentice Hall Flame Test Lab Answers Public Domain eBooks
 - Prentice Hall Flame Test Lab Answers eBook Subscription Services
 - Prentice Hall Flame Test Lab Answers Budget-Friendly Options

6. Navigating Prentice Hall Flame Test Lab Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Prentice Hall Flame Test Lab Answers Compatibility with Devices
 - Prentice Hall Flame Test Lab Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Prentice Hall Flame Test Lab Answers
 - Highlighting and Note-Taking Prentice Hall Flame Test Lab Answers
 - Interactive Elements Prentice Hall Flame Test Lab Answers
8. Staying Engaged with Prentice Hall Flame Test Lab Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Prentice Hall Flame Test Lab Answers
9. Balancing eBooks and Physical Books Prentice Hall Flame Test Lab Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Prentice Hall Flame Test Lab Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Prentice Hall Flame Test Lab Answers
 - Setting Reading Goals Prentice Hall Flame Test Lab Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Prentice Hall Flame Test Lab Answers
 - Fact-Checking eBook Content of Prentice Hall Flame Test Lab Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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