

Chapter 10 - Chemical Quantities

10.1 The Mole: A Measurement of Matter:

Mole: (Thank, Google's, videos)

- Unit of measurement

Representative Particles:

- Smallest unit of a substance (atom, molecule, formula unit, ions, anything...)
- Avogadro's number = 6.02×10^{23} representative particles (RP)
- 6.02×10^{23} representative particles = 1 mol

Molar Mass:

- Molar mass = g/mol
- Molar mass of:
 - an element = atomic mass (g)
Ex. 24.3 g/mol = molar mass of Mg
 - a compound = sum of the atomic masses (g) of each element
Ex. 58.44g/mol = molar mass of NaCl
- Molar mass = 1 mol

Molar Volume:

- Molar volume = 22.4 L/mol
- Standard temperature and pressure (STP) =
 - temp = 0°C
 - pressure = 1 atm
- 22.4 L = 1 mol (at STP)

Calculating Molar mass from Density:

- molar mass (g/mole) = density at STP (g/L) $\times$ molar volume at STP (22.4L/mole)

$$\text{Molar mass (g/mol)} = \left(\frac{\text{grams}}{\text{L}} \right) \left(\frac{22.4 \text{ L}}{\text{mole}} \right)$$

Calculating Density from Molar mass:

- Density at STP (g/L) = molar mass (g/mole) $\times$ molar volume at STP (1mole/22.4L)

$$\text{Density (g/L)} = \left(\frac{\text{grams}}{1 \text{ mole}} \right) \left(\frac{1 \text{ mole}}{22.4 \text{ L}} \right)$$

Prentice Hall Chemistry Chemical Quantities Chapter 10

Lecture Notes

RS Peters

A decorative graphic consisting of a light blue horizontal bar with a rounded right end, and a red circular shape partially visible behind it.

Prentice Hall Chemistry Chemical Quantities Chapter 10 Lecture Notes:

Reviewing **Prentice Hall Chemistry Chemical Quantities Chapter 10 Lecture Notes**: Unlocking the Spellbinding Force of Linguistics

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