#### Worksheet: Acids, Bases, and Salts

Nome KEY

#### What are Acids, Bases, and Salts?

The Swedish chemist Svante Arrhenius introduced the theory of ionization and used this theory to explain much about the behavior of acids and bases.

 $HNO_3(aq) \rightarrow H^*(aq) + NO_3^*(aq)$ 

 $HCI(aq) \rightarrow H'(aq) + CI(aq)$ 

 An Arrhenius base is defined as any compound that dissociates in aqueous solution to form \_\_\_\_\_\_ ions.
 KOH (aq) → K\* (aq) + OH\* (aq)

NaOH (aq)  $\rightarrow$  \_\_\_\_\_ Na^\* (aq) + OH (aq)

Using the Arrhenius definition, classify the following examples as acids, bases, or salts:

HBr asset of KCI Service. beaute act d Mig(OHI)2 HaPO<sub>4</sub> enert of HCIO act d HCL partie. Benzer KNIO<sub>2</sub> Al(OH)s HFO4 act of KC2H3O2 partit. bease Section 1 Ba(OH)<sub>2</sub> NaCl

# **Note Taking Worksheet Acids Bases And Salts**

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