

c. Standard Enthalpy of Formation ΔH_f°

→ The enthalpy change for the formation of 1 mole of liquid or a gas at the boiling point and standard pressure.

Example:



d. Standard Enthalpy Change of Melting ΔH_m°

→ The enthalpy change for melting 1 mol of a solid into a liquid at the melting point and standard pressure.

Example:



e. Standard Neutralization Enthalpy Change ΔH_n°

→ The enthalpy change in the neutralization of an acid (H⁺) by a base (OH⁻) forming 1 mole of water.

Example:



f. Standard Enthalpy of Burning ΔH_b°

→ The enthalpy change for burning 1 mol of elemental atoms in the gas phase at standard conditions.

Example:



g. Standard Solution Enthalpy Change ΔH_{sol}°

→ The enthalpy change when 1 mole of a compound is converted into a solution at standard conditions.

Example:



Note Taking Guide For Thermochemical Equations

M Mark



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Note Taking Guide For Thermochemical Equations Book Review: Unveiling the Power of Words

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