

Parts of the Atom

**Fill in the following table about the parts of the atom. (AMU stands for atomic mass unit.)*

PART OF THE ATOM	ELECTRIC CHARGE	LOCATION IN ATOM	ATOMIC MASS
proton	proton	Nucleus	2000 X More than an e ⁻
neutron	neutron	Nucleus	2000 X More than an e ⁻
electron	electron	Orbitals / Energy Levels	1

**Here's how to predict the structure of any element on the periodic table. For example, look at carbon:*

6
C
carbon
12.011

Error: Reference source not foundatomic number: This is the number of protons (also the number of electrons)

Error: Reference source not foundchemical symbol: This is an abbreviation for the element.

Error: Reference source not foundelement name: This is the name of the element.

Error: Reference source not foundatomic mass: This number (rounded off) tells you the number of protons *plus* the number of neutrons. (Electrons are too small to be included.)

To find the number of protons, look at the atomic number:

Example for carbon:

Atomic number: 6

Number of protons: 6

To find the number of neutrons, subtract the atomic number from the atomic mass:

Example for carbon:

Atomic mass: 12 (rounded)

Number of neutrons: $12 - 6 = 6$

To find the number of electrons, look at the atomic number:

Example for carbon:

Atomic number: 6

Number of electrons: 6

**Complete the table for the elements with atomic numbers 1-10.*

ELEMENT NAME	ATOMIC NUMBER	ATOMIC MASS	PROTONS	NEUTRONS	ELECTRONS
Hydrogen	1	1.00	1	1	1
Helium	2	4.00	2	2	2
Lithium	3	6.91	3	3	3
Beryllium	4	9.01	4	4	4
Boron	5	10.81	5	5	5
Carbon	6	12.01	6	6	6
Nitrogen	7	14.00	7	7	7
Oxygen	8	15.99	8	8	8

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B Lingard

A red circular graphic with a gradient, appearing as a stylized arrow or a partial circle, located to the right of the name B Lingard.

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