

Equation Writing and Predicting Products

Introduction:

If you examine a car that has been in a junkyard for a while, you will notice that it has rusted. Rusting is a slow chemical reaction of the iron in the car's body with oxygen gas. If a piece of sodium is put into water, a much more rapid chemical reaction occurs. Sodium reacts with water to produce sodium hydroxide and hydrogen gas. During this reaction, enough heat is liberated to ignite the hydrogen gas, causing it to explode.

Chemists observe what is happening in a chemical reaction and try to describe it in language that is simple and clear. A chemical equation uses formulas and symbols to describe the substances involved in a reaction, the physical state of the substance, the use of a catalyst, and relative proportions. The general form of an equation is:



In this investigation you will perform a series of reactions and make careful observations of the changes that occur. Using simple tests and your knowledge of chemistry, you will determine the identity of the products. With this information, you will write chemical equations to describe the reactions.

Pre-Lab Discussion:

Read through the entire lab investigation and the relevant pages of your textbook. Then answer the questions that follow.

1. What constitutes a positive test for each of the following gases?
 - a. oxygen _____
 - b. hydrogen _____
 - c. carbon dioxide _____
 - d. water vapor _____
 - e. ammonia (NH_3) _____
2. What is the proper way to smell a substance in the lab? Why should care be taken when smelling a gas such as ammonia? _____

3. What is the role of a catalyst in a reaction? How can you tell when a substance serves as a catalyst? (This answer will be important later, spend some time on it now.) _____

4. One way to identify limestone (CaCO_3) is to drop a small amount of hydrochloric acid (HCl) on it. A positive test results in a fire of carbon dioxide being produced. Write out in words the information represented by the following balanced chemical equation.
$$\text{CaCO}_3(s) + 2\text{HCl}(aq) \longrightarrow \text{CaCl}_2(aq) + \text{H}_2\text{O}(l) + \text{CO}_2(g)$$

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James F. Hall



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